

Objective: Students will identify and differentiate between the different parts and models of the atom.

1. Atoms are made of three basic particles: _____
2. What is electric charge? _____
3. A positive and a negative charge will _____
4. Two positive charges will _____
5. Two negative charges will _____
6. What does the term *isotope* represent? _____
7. Electric charge only appears in _____ of the elementary charge.
8. What did J.J. Thompson discover? _____
9. Atoms have _____ charge.
10. The model thought electrons were sprinkled around inside atoms like _____ in a _____ of _____.
11. This was the _____ ball model for the _____ of an _____, but it wasn't the _____ model.
12. In 1911 Rutherford, Geiger and Marsden did an experiment where they shot helium at a thin gold foil. They expected most of the helium to be deflected a little. What actually happened? _____

13. The best way to explain this was if a gold atom was mostly _____.
14. How much of an atom's mass is contained in the nucleus? _____
15. Where is the electron found? _____
16. Where is the proton found? _____
17. Where is the neutron found? _____
18. How much bigger is the electron cloud than the nucleus? _____
19. If the atom were a football stadium, how big would the nucleus be? _____
How big would the electron be? _____
20. What is the charge of a proton? _____
21. What is the charge of an electron? _____
22. What is the charge of a neutron? _____
23. Why doesn't an electron fall into the nucleus? _____
24. What holds the nucleus together? _____
25. What are the four types of forces inside an atom? _____

Foundations Of Physical Science Answers To Chapter 9

Clifford Lane



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