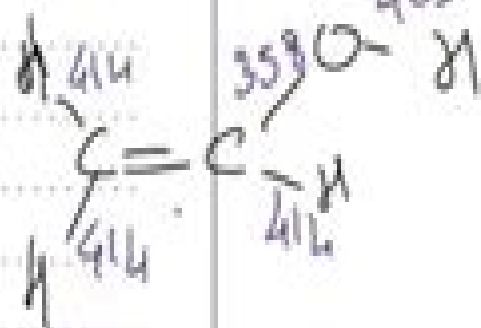
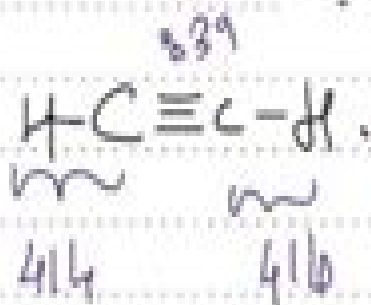


- (ii) Determine the enthalpy change for the reaction, in kJ, to produce A using section 11 of the data booklet.

[3]



$$\text{C-H: } 414 \times 2$$

$$\text{OH: } 463 \times 2$$

$$414 \times 3 \rightarrow \text{CH}$$

$$\text{C}\equiv\text{C} \quad 839$$

$$\begin{array}{cc} 358 & \text{CO} \\ 463 & \text{OH} \end{array}$$

$$\Delta H_2 = E_R - E_P$$

$$= 414 \times 2$$

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