

Suggested Answers

VCE Chemistry 2016 Year 11 Trial Exam Unit 2

SECTION A – Multiple Choice Answers

(1 mark per question)

- Q1. C** Referring to the *VCE Chemistry Data Book: Table 3*
Specific heat capacity for water: $c = 4.18 \text{ J g}^{-1} \text{ }^{\circ}\text{C}^{-1}$
Density of water = 1.00 g mL^{-1}
 $5.00 \text{ L} = 5.00 \times 1000 = 5000 \text{ mL} = 5000 \text{ g}$
 $\Delta T = 95.0 - 25.0 = 70.0 \text{ }^{\circ}\text{C}$
 $Q = m \times c \times \Delta T = 5000 \times 4.18 \times 70.0 = 1.46 \times 10^6 \text{ J}$
 $Q = 1.46 \times 10^6 / 1000 = 1460 \text{ kJ}$
- Q2. B** Like hydrogen sulfide, water is a Group 16 hydride.
The boiling point of hydrogen sulfide is significantly lower than that for water ($-60 \text{ }^{\circ}\text{C}$ compared to $100 \text{ }^{\circ}\text{C}$).
Since the boiling temperature reflects the bonding interactions **between** the molecules, the forces attracting the water molecules to each other are greater. Both water and hydrogen sulfide are polar molecules because of the difference in the electronegativities of hydrogen, oxygen and sulfur. Therefore there will be dipole-dipole interactions between both H_2O and H_2S molecules. However, in water there are stronger hydrogen bonding interactions between the molecules and this accounts for the increased temperature of the boiling point.
- Q3. C** All common nitrate compounds are soluble in water.
Chlorides other than those of silver and lead(II) are soluble.
Therefore, when aqueous solutions of cobalt(II) chloride and silver nitrate are mixed, a precipitate of silver chloride will form.
Full equation:
 $\text{CoCl}_2(\text{aq}) + 2\text{AgNO}_3(\text{aq}) \rightarrow \text{Co}(\text{NO}_3)_2(\text{aq}) + 2\text{AgCl}(\text{s})$
Ionic equation:
 $\text{Cl}^-(\text{aq}) + \text{Ag}^+(\text{aq}) \rightarrow \text{AgCl}(\text{s})$

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