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Chapter 1 Introduction

1.1 A gas at 20°C may be rangled if it contains less than 10¹² molecules per mm³. If Avogadro's number is 6.023E23 molecules per mole, what air pressure does this represent?

Solution: The mass of one molecule of air may be computed as

Then the density of air containing 1012 molecules per mm3 is, in SI units,

$$\frac{g}{mm^3} = \frac{12 \text{ molecules}}{mm^3} = 4.81E-5 \cdot \frac{g}{m}$$
= 4.81E-11 \frac{g}{mm^3} = 4.81E-5 \frac{kg}{m}_3

Finally, from the perfect gas law, Eq. (1.13), at 20°C = 293 K, we obtain the pressure:

$$p = \rho RT = 4.8185 - 5$$
 $\frac{\text{kg}}{\text{m}^3} = \frac{\text{m}^2}{\text{s}^2 \cdot \text{K}}$ (293 K) = 4.0 Pa Ans.

1.2 The earth's atmosphere can be modeled as a uniform layer of air of thickness 20 km and average density 0.6 kg/m³ (see Table A-6). Use these values to estimate the total mass and total number of molecules of air in the entire atmosphere of the earth.

Solution: Let R_e be the earth's radius ≈ 6377 km. Then the total mass of air in the atmosphere is

$$m_t = -\rho \, d\text{Vol} = \rho_{avg} \, (\text{Air Vol}) \approx \rho_{avg} \, 4\pi \, \text{R}^2_{c} \, (\text{Air thickness})$$

= $(0.6 \, \text{kg/m}^3) 4\pi \, (6.377E6 \, \text{m})^2 \, (20E3 \, \text{m}) \approx 6.1E18 \, \text{kg}$ Ans.

Dividing by the mass of one molecule ≈ 4.8E-23 g (see Prob. 1.1 above), we obtain the total number of molecules in the earth's atmosphere:

$$\frac{N}{motorules} = \frac{m(atmosphere)}{m(one molecule)} = \frac{6.1E21 \text{ grams}}{4.8E-23 \text{ gm/molecule}} \approx 1.3E44 \text{ molecules}$$
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Robert Lynnwood Street, John K. Vennard, Bruce M. Howe

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