



Nov 9, 2015, 9:41 AM



Name: Ali Washington

Date: 11/9/15

Class: 9

5 INTERPRETING GRAPHICS

Use with Section 5.3



Figure 1 The emission spectrum and orbit-transition diagram for hydrogen.

speed of light
 $c = 3 \times 10^8 \text{ m/s}$
 $c = \lambda \cdot \nu$ (wave length)
 $E = h \cdot \nu$ (energy)
 $h = 6.626 \times 10^{-34} \text{ Js}$
 Planck's constant

1) $2.46 \times 10^{-18} = 6.626 \times 10^{-34} \nu$
 $\nu = \frac{2.46 \times 10^{-18}}{6.626 \times 10^{-34}} = 3.71 \times 10^{15} \text{ s}^{-1}$
 $\lambda = \frac{c}{\nu} = \frac{3 \times 10^8}{3.71 \times 10^{15}} = 8.11 \times 10^{-8} \text{ m}$

Transition	Energy, J (J)	ν (s ⁻¹)	λ (nm)	Type of Radiation
$n=5 \rightarrow n=3$	2.66×10^{-19}	4.01×10^{15}	7.48×10^{-8}	IR
$n=5 \rightarrow n=4$	7.57×10^{-20}	1.14×10^{16}	2.63×10^{-8}	IR
$n=5 \rightarrow n=3$	1.82×10^{-19}	1.4×10^{16}	2.13×10^{-8}	IR
$n=5 \rightarrow n=2$	4.84×10^{-19}	6.76×10^{16}	4.44×10^{-9}	visible
$n=5 \rightarrow n=1$	2.82×10^{-18}	3.20×10^{16}	9.36×10^{-9}	UV
$n=5 \rightarrow n=4$	4.90×10^{-20}	7.40×10^{15}	4.01×10^{-8}	IR
$n=5 \rightarrow n=3$	1.55×10^{-19}	2.35×10^{16}	1.28×10^{-8}	IR
$n=5 \rightarrow n=2$	4.56×10^{-19}	6.91×10^{16}	4.34×10^{-9}	visible
$n=5 \rightarrow n=1$	2.69×10^{-18}	3.17×10^{16}	9.46×10^{-9}	UV
$n=4 \rightarrow n=3$	1.06×10^{-19}	1.61×10^{16}	1.86×10^{-8}	IR
$n=4 \rightarrow n=2$	4.09×10^{-19}	6.17×10^{16}	4.99×10^{-9}	visible
$n=4 \rightarrow n=1$	2.04×10^{-18}	3.08×10^{16}	9.74×10^{-9}	UV
$n=3 \rightarrow n=2$	3.03×10^{-19}	4.57×10^{16}	6.56×10^{-9}	visible
$n=3 \rightarrow n=1$	1.94×10^{-18}	2.93×10^{16}	1.21×10^{-8}	UV
$n=2 \rightarrow n=1$	1.64×10^{-18}	2.47×10^{16}	1.21×10^{-8}	UV

$h = 6.626 \times 10^{-34} \text{ J s}$

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A Loxley

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