

## 18. Lattice enthalpies at 298 K (experimental values)

The lattice enthalpy values ( $\Delta H_{\text{lattice}}^\circ$ ) given relate to the endothermic process  $M_aX_b(s) \rightarrow aM^{b+}(g) + bX^{a-}(g)$  in which the gaseous ions of a crystal are separated to an infinite distance from each other.

### Experimental values

The data in these tables are experimental values obtained by means of a suitable Born-Haber cycle.

Alkali metal halides	$\Delta H_{\text{lattice}}^\circ$ (kJ mol <sup>-1</sup> )			
	F	Cl	Br	I
Li	1049	864	820	764
Na	930	790	754	705
K	829	720	691	650
Rb	795	695	668	632
Cs	759	670	647	613

Other substances	$\Delta H_{\text{lattice}}^\circ$ (kJ mol <sup>-1</sup> )
CaF <sub>2</sub>	2651
BeCl <sub>2</sub>	3033
MgCl <sub>2</sub>	2540
CaCl <sub>2</sub>	2271
SrCl <sub>2</sub>	2170
BaCl <sub>2</sub>	2069
MgO	3791
CaO	3401

Other substances	$\Delta H_{\text{lattice}}^\circ$ (kJ mol <sup>-1</sup> )
SrO	3223
BaO	3054
CuCl <sub>2</sub>	2824
AgF	974
AgCl	918
AgBr	905
AgI	892

# Ib Chemistry Data Let

**SJ Ball**



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