

Question 33 (4 marks)

A student adds 1.17 g of $\text{Al}(\text{OH})_3(\text{s})$ to 0.500 L of $0.100 \text{ mol L}^{-1} \text{HCl}(\text{aq})$.

4

Calculate the pH of the resulting solution. Assume that the volume of the resulting solution is 0.500 L.

① Write a balanced chemical equation



② Find no. of moles of OH^- and H^+

$$n(\text{Al}(\text{OH})_3) = \frac{m}{M_r} = \frac{1.17}{26.98 + 3(16 + 1.008)} = 0.04499 \dots \text{mol}$$

$$n(\text{OH}^-) = 3 \times n(\text{Al}(\text{OH})_3) = 0.13497 \dots \text{mol}$$

$$n(\text{HCl}) = c \times v = 0.5 \times 0.1 = 0.05 \text{ mol}$$

$$n(\text{H}^+) = n(\text{HCl}) = 0.05 \text{ mol}$$

③ Determine which reagent is in excess and how much

As $n(\text{H}^+) > n(\text{OH}^-)$ $\therefore \text{H}^+$ is in excess

$$n(\text{H}^+)_{\text{excess}} = 0.05 - 0.13499 \dots = 0.00500 \dots \text{mol}$$

④ Calculate the $[\text{H}^+]$ and thus find pH.

$$[\text{H}^+] = \frac{n}{V} = \frac{0.00500}{0.500} = 0.0100 \text{ mol L}^{-1}$$

$$\text{pH} = -\log_{10} [\text{H}^+] = 2.000 \text{ (3 dp)}$$

Using d.p. 0015 for pH

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Hsc Chemistry 6 User Guide

Pierre Steinmetz



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Table of Contents Hsc Chemistry 6 User Guide

1. Understanding the eBook Hsc Chemistry 6 User Guide
 - The Rise of Digital Reading Hsc Chemistry 6 User Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Chemistry 6 User Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Chemistry 6 User Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Chemistry 6 User Guide
 - Personalized Recommendations
 - Hsc Chemistry 6 User Guide User Reviews and Ratings
 - Hsc Chemistry 6 User Guide and Bestseller Lists
5. Accessing Hsc Chemistry 6 User Guide Free and Paid eBooks
 - Hsc Chemistry 6 User Guide Public Domain eBooks
 - Hsc Chemistry 6 User Guide eBook Subscription Services
 - Hsc Chemistry 6 User Guide Budget-Friendly Options

6. Navigating Hsc Chemistry 6 User Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Hsc Chemistry 6 User Guide Compatibility with Devices
 - Hsc Chemistry 6 User Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Chemistry 6 User Guide
 - Highlighting and Note-Taking Hsc Chemistry 6 User Guide
 - Interactive Elements Hsc Chemistry 6 User Guide
8. Staying Engaged with Hsc Chemistry 6 User Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Chemistry 6 User Guide
9. Balancing eBooks and Physical Books Hsc Chemistry 6 User Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Chemistry 6 User Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Hsc Chemistry 6 User Guide
 - Setting Reading Goals Hsc Chemistry 6 User Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Hsc Chemistry 6 User Guide
 - Fact-Checking eBook Content of Hsc Chemistry 6 User Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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