



**55. Phosphorus has three allotropic forms — (i) white phosphorus (ii) red phosphorus and (iii) black phosphorus. Write the difference between white and red phosphorus on the basis of their structure and reactivity.**

**Solution:**

1. White phosphorus

(i) It is soft waxy solid having garlic order

(ii) It is poisonous

(iii) Has low melting point and boiling point because P<sub>4</sub> molecules are held together by weak van der Waal forces of attraction

2. Red Phosphorus

(i) It is hard, crystalline, odourless solid

(ii) It is non-poisonous

(iii) It has a high melting point because of its polymeric structure.

**56. Give an example to show the effect of concentration of nitric acid on the formation of oxidation product.**

**Solution:**



**57. PCI<sub>5</sub> reacts with finely divided silver on heating and a white silver salt is obtained, which dissolves on adding excess aqueous NH<sub>3</sub> solution. Write the reactions involved to explain what happens.**

**Solution:**

PCI<sub>5</sub> reacts with silver to form white silver salt(AgCl). Which then dissolves on adding excess aqueous NH<sub>3</sub> to form a soluble complex



**58. Phosphorus forms a number of oxoacids. Out of these oxoacids phosphinic acid has strong**

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1. Understanding the eBook Glencoe Chemistry Chapter Answers
  - The Rise of Digital Reading Glencoe Chemistry Chapter Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Glencoe Chemistry Chapter Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Glencoe Chemistry Chapter Answers
  - User-Friendly Interface

4. Exploring eBook Recommendations from Glencoe Chemistry Chapter Answers
  - Personalized Recommendations
  - Glencoe Chemistry Chapter Answers User Reviews and Ratings
  - Glencoe Chemistry Chapter Answers and Bestseller Lists
5. Accessing Glencoe Chemistry Chapter Answers Free and Paid eBooks
  - Glencoe Chemistry Chapter Answers Public Domain eBooks
  - Glencoe Chemistry Chapter Answers eBook Subscription Services
  - Glencoe Chemistry Chapter Answers Budget-Friendly Options
6. Navigating Glencoe Chemistry Chapter Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Glencoe Chemistry Chapter Answers Compatibility with Devices
  - Glencoe Chemistry Chapter Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Glencoe Chemistry Chapter Answers
  - Highlighting and Note-Taking Glencoe Chemistry Chapter Answers
  - Interactive Elements Glencoe Chemistry Chapter Answers
8. Staying Engaged with Glencoe Chemistry Chapter Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Glencoe Chemistry Chapter Answers
9. Balancing eBooks and Physical Books Glencoe Chemistry Chapter Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Glencoe Chemistry Chapter Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Glencoe Chemistry Chapter Answers
  - Setting Reading Goals Glencoe Chemistry Chapter Answers
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Glencoe Chemistry Chapter Answers
  - Fact-Checking eBook Content of Glencoe Chemistry Chapter Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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