

- and used in surgical instrument
4. Tin lowers the melting point of the alloy and is used for soldering purpose. D. Stainless steel
5. The alloy is hard, brittle, takes up polish and is used for making statues E. Solder

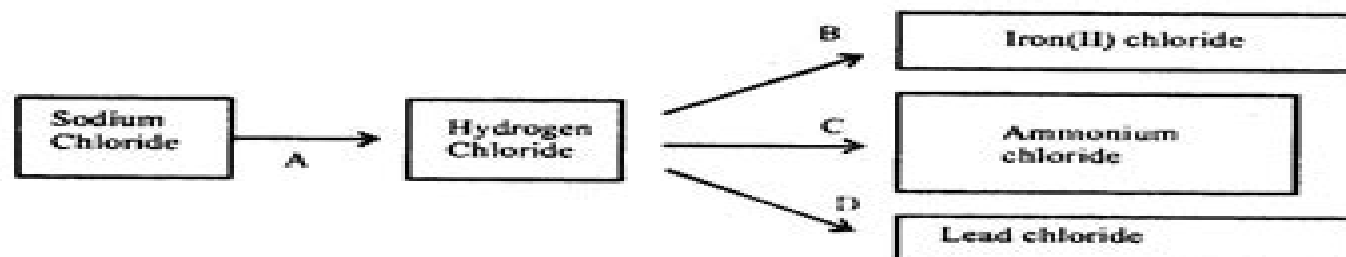
[5]

Question 4

- (a) Identify the anion present in the following compounds:
- Compound X on heating with copper turnings and concentrated sulphuric acid liberates a reddish brown gas.
 - When a solution of compound Y is treated with silver nitrate solution a white precipitate is obtained which is soluble in excess of ammonium hydroxide solution.
 - Compound Z which on reacting with dilute sulphuric acid liberates a gas which turns lime water milky, but the gas has no effect on acidified potassium dichromate solution.
 - Compound L on reacting with Barium chloride solution gives a white precipitate insoluble in dilute hydrochloric acid or dilute nitric acid. [4]
- (b) State one chemical test between each of the following pairs:
- Sodium carbonate and Sodium sulphite
 - Ferrous nitrate and Lead nitrate
 - Manganese dioxide and Copper (II) oxide [3]
- (c) Draw an electron dot diagram to show the structure of hydronium ion. State the type of bonding present in it. [3]

Question 5

- (a) (i) 67.2 litres of hydrogen combines with 44.8 litres of nitrogen to form ammonia under specific conditions as:
- $$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \longrightarrow 2\text{NH}_3(\text{g})$$
- Calculate the volume of ammonia produced. What is the other substance, if any, that remains in the resultant mixture? [2]
- (ii) The mass of 5.6 dm³ of a certain gas at STP is 12.0 g. Calculate the relative molecular mass of the gas. [2]
- (iii) Find the total percentage of Magnesium in magnesium nitrate crystals, Mg(NO₃)₂·6H₂O. [Mg = 24; N = 14; O = 16 and H = 1] [2]
- (b) Refer to the flow chart diagram below and give balanced equations with conditions, if any, for the following conversions A to D. [4]



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Table of Contents Icse Class 10 Question Papers 2012

1. Understanding the eBook Icse Class 10 Question Papers 2012
 - The Rise of Digital Reading Icse Class 10 Question Papers 2012
 - Advantages of eBooks Over Traditional Books
2. Identifying Icse Class 10 Question Papers 2012
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Icse Class 10 Question Papers 2012
 - User-Friendly Interface
4. Exploring eBook Recommendations from Icse Class 10 Question Papers 2012
 - Personalized Recommendations
 - Icse Class 10 Question Papers 2012 User Reviews and Ratings

- Icse Class 10 Question Papers 2012 and Bestseller Lists
- 5. Accessing Icse Class 10 Question Papers 2012 Free and Paid eBooks
 - Icse Class 10 Question Papers 2012 Public Domain eBooks
 - Icse Class 10 Question Papers 2012 eBook Subscription Services
 - Icse Class 10 Question Papers 2012 Budget-Friendly Options
- 6. Navigating Icse Class 10 Question Papers 2012 eBook Formats
 - ePub, PDF, MOBI, and More
 - Icse Class 10 Question Papers 2012 Compatibility with Devices
 - Icse Class 10 Question Papers 2012 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Icse Class 10 Question Papers 2012
 - Highlighting and Note-Taking Icse Class 10 Question Papers 2012
 - Interactive Elements Icse Class 10 Question Papers 2012
- 8. Staying Engaged with Icse Class 10 Question Papers 2012
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Icse Class 10 Question Papers 2012
- 9. Balancing eBooks and Physical Books Icse Class 10 Question Papers 2012
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Icse Class 10 Question Papers 2012
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Icse Class 10 Question Papers 2012
 - Setting Reading Goals Icse Class 10 Question Papers 2012
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Icse Class 10 Question Papers 2012
 - Fact-Checking eBook Content of Icse Class 10 Question Papers 2012
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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