Problem Solving continued

12. Urea, $(NH_2)_2CO$, is an important fertilizer that is manufactured by the following reaction.

$$2NH_3(g) + CO_2(g) \rightarrow (NH_2)_2CO(s) + H_2O(g)$$

What volume of NH_3 at STP will be needed to produce 8.00×10^4 kg of urea if there is an 89.5% yield in the process?

13. An obsolete method of generating oxygen in the laboratory involves the decomposition of barium peroxide by the following equation.

$$2BaO_2(s) \rightarrow 2BaO(s) + O_2(g)$$

What mass of BaO_2 reacted if 365 mL of O_2 is collected by water displacement at 0.975 atm and 10.°C?

14. It is possible to generate chlorine gas by dripping concentrated HCl solution onto solid potassium permanganate according to the following equation.

$$2KMnO_4(s) + 16HCl(aq) \rightarrow 2KCl(aq) + 2MnCl_2(aq) + 8H_2O(l) + 5Cl_2(g)$$

If excess HCl is dripped onto 15.0 g of $KMnO_4$, what volume of Cl_2 will be produced? The Cl_2 is measured at 15°C and 0.959 atm.

15. Ammonia can be oxidized in the presence of a platinum catalyst according to the following equation.

$$4NH_3(g) + 5O_2(g) \rightarrow 4NO(g) + 6H_2O(l)$$

The NO that is produced reacts almost immediately with additional oxygen according to the following equation.

$$2NO(g) + O_2(g) \rightarrow 2NO_2(g)$$

If 35.0 kL of oxygen at STP react in the first reaction, what volume of NH_3 at STP reacts with it? What volume of NO_2 at STP will be formed in the second reaction, assuming there is excess oxygen that was not used up in the first reaction?

16. Oxygen can be generated in the laboratory by heating potassium chlorate. The reaction is represented by the following equation.

$$2KClO_3(s) \rightarrow 2KCl(s) + 3O_2(g)$$

What mass of $KClO_3$ must be used in order to generate 5.00 L of O_2, measured at STP?

17. One of the reactions in the Solvay process is used to make sodium hydrogen carbonate. It occurs when carbon dioxide and ammonia are passed through concentrated salt brine. The following equation represents the reaction.

$$NaCl(aq) + H_2O(l) + CO_2(g) + NH_3(g) \rightarrow NaHCO_3(s) + NH_4Cl(aq)$$

a. What volume of NH_3 at 25°C and 1.00 atm pressure will be required if 38 000 L of CO_2, measured under the same conditions, react to form $NaHCO_3$?

b. What mass of $NaHCO_3$ can be formed when the gases in (a) react with NaCl?

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