



I. DALTON'S ATOMIC THEORY

- John Dalton proposed an atomic theory of matter that can explain chemical observation as predicted by the fundamental laws:

1. **Law of Conservation** states that matter is not created nor destroyed.
2. **Law of definite proportion** states that different samples of the same compound always contain the constituent elements in the same proportion by mass.
3. **Law of multiple proportion** states that if two elements form more than one compound, the various masses of one element combining with a fixed mass of another element are related by small whole-number ratios.

ATOMIC THEORY

1. Atoms are the smallest particles of matter. They cannot be divided into smaller particles, they also cannot be created nor destroyed.
2. All atoms of an element are identical, but the atoms of one element are different from atoms of other elements in terms of mass, size, and properties.
3. Compounds are composed of atoms of more than one element which are combined in fixed ratios.
4. Atoms retain their identity during chemical reactions, which involve combination, separation, and rearrangement. They are indestructible.

Matter And Atoms Study Guide

Kathleen Armour



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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Matter And Atoms Study Guide**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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