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Acid-Base Theories

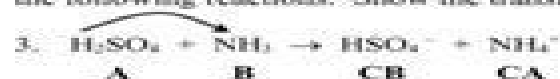
1. Summarize the three main acid-base theories in the table below.

	ACID	BASE
Arrhenius	H^+ or H_3O^+ producer in aqueous solution	OH^- producer in aqueous solution
Bronsted-Lowry	proton (H^+) donor	proton (H^+) acceptor
Lewis	electron-pair acceptor	electron-pair donor

Use the following chemical equations to answer question 2.



2. According to Arrhenius, HCl would be classified as a(n)
- acid
- and
- NH_3
- would be classified as a(n)
- base
- . Explain why.

HCl produces hydronium ions (H_3O^+) in water. NH_3 produces hydroxide ions (OH^-) in water.For problems 3 – 5, label the acid (A), base (B), conjugate acid (CA), and conjugate base (CB) in each of the following reactions. Show the transfer of the proton (H^+) by drawing an arrow.

For problems 6 – 9, give the conjugate base for each Bronsted-Lowry acid.



For problems 10 – 13, give the conjugate acid for each Bronsted-Lowry base.



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