Section 7.2: Equilibrium Law and the Equilibrium Constant Tutorial 1 Practice, page 431

1. (a) $2 \text{ CO}_2(g) \implies 2 \text{ CO}(g) + O_2(g)$ Products: 2 CO(g); O₂(g) Reactant: 2 CO₂(g) Equilibrium law equation: $K = \frac{[CO(g)]^2[O_2(g)]}{[CO_2(g)]^2}$ (b) $2 \text{ Cl}_2(g) + 2 \text{ H}_2O(g) \implies 4 \text{ HCl}(g) + O_2(g)$ Products: 4 HCl(g); O₂(g) Reactants: 2 Cl₂(g); 2 H₂O(g) Equilibrium law equation: $K = \frac{[HCl(g)]^4[O_2(g)]}{[Cl_2(g)]^2[H_2O(g)]^2}$ (c) $2 O_0(g) \implies 3 O_2(g)$ Product: $3 O_2(g)$ Reactant: 2 O_y(g)

Equilibrium law equation: $K = \frac{[O_2(g)]^3}{[O_2(g)]^2}$

(d) $4 \text{ NH}_3(g) + 3 \text{ O}_2(g) \implies 2 \text{ N}_2(g) + 6 \text{ H}_2\text{O}(g)$ Products: $2 N_2(g)$; $6 H_2O(g)$ Reactants: $4 \text{ NH}_3(g)$; $3 \text{ O}_2(g)$

Equilibrium law equation: $K = \frac{[N_2(g)]^2[H_2O(g)]^6}{[NH_2(g)]^4[O_2(g)]^3}$

 Given: CO(g) + 2 H₂(g) = CH₃OH(g); [CO(g)]_{equilibrium} = 0.079 mol/L; $[H_2(g)]_{\text{opathteium}} = 0.158 \text{ mol/L}; [CH_2OH(g)]_{\text{opathteium}} = 0.021 \text{ mol/L}.$ Required: K

Solution: Write the equilibrium law equation using the balanced chemical equation. Then, substitute the equilibrium concentrations into the equilibrium law equation and solve for K.

$$K = \frac{[CH_3OH(g)]}{[CO(g)][H_2(g)]^2}$$

$$= \frac{(0.021)}{(0.079)(0.158)^2}$$

K = 11

Statement: The equilibrium constant, K, for the reaction is 11.

Grade 12 Nelson Chemistry Questions Answers

A Loxley

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