CHEMICAL KINETICS SAMPLE REPORT

Special Notes

- The report must be written in ink (blue pen), no pencil allowed. However, the graph can be drawn in pencil.
- The layout of the report must be tidy (no white out) and legible in writing.(Work out all the solutions on some scrap papers first, then copy to the report sheets.)
- Page 7 1
 - Record exact temperatures in °C to <u>1 decimal</u>.
 - ② Both the "Time" in seconds and "Relative Rate" must be integers.
 - Report the "Initial [I]", "Initial [S₂O₈²]" and "Rate" in 3 s.f.
- Page 7 2
 - Show all calculations for:
 - (A) Concentration of initial [I"],
 - (B) Concentration of initial [S₂O₈²⁻],
 - (C) Rate of Reaction of S₂O_x²⁻, and
 - (D) Relative Rates
 - Keep all calculations in 3 s.f.. However, Relative Rate, Order of Reaction, and "m" and "n" must be reported as integers.
- Page 7 3
 - ① For Question I, the conclusion must include: rate ~ [I-].
 - To Proposition 2, the conclusion must include: rate = $[S_2O_8^{2-}]$.
 - \odot For Question 3, the rate law is: rate = $k[I^-][S_2O_8^{2-}]$
- Page 7 4
 - Show calculations of rate constants, k, for all five reactions.
 - ② The <u>units</u> of rate constant, k, is M⁻¹-s⁻¹.
 - ① The accuracy of k is between 3.30×10^{-3} to 3.90×10^{-3} M⁻¹-s⁻¹ at 22-23°C.
 - The precision of k ±0.30 × 10⁻3 M⁻¹⁻s⁻¹ between highest and lowest k.

Example Lab Report Kinetics Of A Reaction

JS Bruner

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