

A chemical bond - forms when 2 or more atoms rearrange valence  
electrons to increase stability.

ionic bond - forms when valence electrons are transferred  
from one atom to another

cation - atom loses electrons to become + charged

anion - atom gains electrons to become - charged

In ionic compounds the ions are arranged in a crystal lattice.  
Strong forces hold the ions together.

Properties of ionic compounds:

- high melting and boiling points
- hard - not easily crushed
- conduct electricity when melted or  
dissolved because the ions are free to move.

covalent bond - electrons are shared, forming  
molecules.

Covalent compounds have weaker forces holding the  
molecules together.

Properties of covalent compounds:

- lower melting and boiling points
- Many covalent compounds are volatile liquids or gases.
- soft - easier to crush
- are not conductors of electricity

electronegativity - property that tells how strong an atom's attraction  
is for electrons.

Since oxygen has a higher electronegativity than hydrogen, oxygen  
holds onto shared electrons tighter, giving the oxygen a partial  
negative charge and the hydrogen a partial positive charge.

# Ionic Bonds And Notetaking Guide Answer

**Michal Rosen-Zvi**



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