

Chapter 6 - Lecture Worksheet 3

Using standard heats of formation, calculate the standard enthalpy change for the following reaction.



$$\begin{aligned}\Delta H^\circ_{\text{rxn}} &= \Delta H^\circ_f(\text{CH}_4, \text{g}) + 2\Delta H^\circ_f(\text{H}_2\text{O}, \text{g}) - \Delta H^\circ_f(\text{CO}_2, \text{g}) - 4\Delta H^\circ_f(\text{H}_2, \text{g}) \\ &= -75 + 2(-242) - (-394) - 4(0) \text{ kJ} = -165 \text{ kJ}\end{aligned}$$

(1) -711 kJ (2) -165 kJ (3) -15 kJ (4) 15 kJ (5) 165 kJ (6) 229 kJ (7) 711 kJ

Using average bond enthalpies estimate the standard enthalpy change for the same reaction.



$$\begin{aligned}\Delta H^\circ_{\text{rxn}} &= 2BE(\text{C}=\text{O}) + 4BE(\text{H}-\text{H}) - 4BE(\text{C}-\text{H}) - 4BE(\text{O}-\text{H}) \\ &\approx 2(728) + 4(436) - 4(413) - 4(463) \text{ kJ} = -304 \text{ kJ}\end{aligned}$$

Which is more accurate ? Why ?

ΔH° calculated from ΔH°_f is more accurate because ΔH° calculated from bond enthalpies uses AVERAGE bond enthalpies obtained from a number of different compounds in which the bonds arise and are not specific to the compound in question.

	<u>Standard Enthalpy of formation (kJ/mol)</u>
$\text{CH}_4(\text{g})$	-75
$\text{CO}_2(\text{g})$	-394
$\text{C}_2\text{H}_2(\text{g})$	+227
$\text{H}_2\text{O}(\text{g})$	-242
$\text{H}_2\text{O}(\text{l})$	-286
$\text{NH}_3(\text{g})$	-46
$\text{HCl}(\text{g})$	-92
$\text{HCl}(\text{aq})$	-167
$\text{CaCO}_3(\text{s})$	-1208
$\text{Ca}(\text{OH})_2(\text{aq})$	-1003
$\text{Ca}(\text{OH})_2(\text{s})$	-986
$\text{CaO}(\text{s})$	-635
$\text{SO}_2(\text{g})$	-297
$\text{SO}_3(\text{g})$	-396

	<u>Average Bond Enthalpies (kJ/mol)</u>
H-H	436
N-H	391
N-N	161
N=N	418
N#N	945
O-H	463
C-O	351
C=O	728
CaO	1072
Cl-Cl	242
Cl-H	432
Br-Br	193
Br-H	366
C-H	413

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