Reaction Rates and Chemical Equilibrium

A. Factors That Affect the Rate of a Reaction

A.1 Effect of Temperature

| Test Tube | Temperature (3) | Observations (4) | | |
|--------------|--------------------|--------------------------|--|--|
| 1. | 6.8°C | Some bubbling | | |
| 2. | 57.4°C | Many bubbles and fizzing | | |

Which test tube cleared first (4)?

Q1 How did an increase in temperature affect the rate of a reaction?

As temperature increases, rate of reaction increases. An increase in temperature will raise average kinetic energy of the reactant molecules. Therefore, a greater proportic molecules will have the minimum energy necessary for an effective collision and maximum energy necessary for an effective collision and maximum energy necessary.

A.2 Effect of Changing Reactant Concentration

| Test Tube | Concentration of HC1 | Initial time (2) | Final time (3) | Total time (3) – (2) |
|--------------|----------------------|---------------------|-------------------|-------------------------|
| L: | 1.0 M HC1 | 0 | | 206 s |
| 2. | 2.0 M HC1 | 0 | | |
| 3. | 3.0 M HC1 | 0 | | 113 s |
| | | | | 30 s |

Lab Report On Reaction Rates And Equilibrium

M Walker

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