	Get the Gizmo ready:
Activity A:	Click Reset (2).
Alpha decay	<ul> <li>Check that Alpha decay and Uranium are selected.</li> </ul>



#### Question: How does alpha decay change the nucleus of a radioactive atom?

- Predict: As you observed in the warm-up activity, an alpha particle consists of two protons and two neutrons. How will the emission of an alpha particle affect the following?
  - A. The atomic number of the atom:

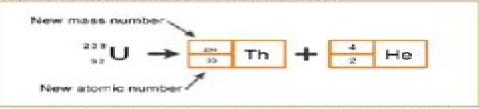
decrease by two

B. The mass number of the atom:

decrease by four

- Calculate: Turn on Write equation. What you see is an equation that shows the original uranium atom on the left. The boxes on the right represent the daughter product—the atom produced by radioactive decay—and the emitted alpha particle.
  - A. In the top left box, write the mass number of the daughter product and press "Enter" on your keyboard. What is this number?

    234
  - B. In the bottom left box, write the atomic number of the daughter product and press "Enter." What is this number?
  - C. \( \) Click on the picture below, select \( \) edit. Enter the mass number and atomic number of the alpha particle, which has the same composition as the nucleus of a helium (He) atom. After filling in the boxes in the Gizmo, write the completed equation below:



D. According to your equation, what isotope remains after the alpha decay of uranium-238? (Note: You can look up element symbols on the periodic table.)

Th - 234

Check: Turn on Show equation, and click Play. The equation will appear at the end of the animation. Was your prediction correct?
 Yes
 If not, modify your equation above.

# **Key Terms Nuclear Changes Answer Key**

**Sabine Zange** 

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