

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Identify a conjugate acid-base pair in chemical reactions involving acids and bases.
2. Calculate the pH of a solution from the concentration of H^+ or H_3O^+ ions, and vice versa.
3. Determine the pH of a solution from K_a and K_b .

Vocabulary

acid	1. conjugate base	acid solution
acid-base indicator	2. hydronium	acidic solution
acidic solution	3. autoionization	pH

Key Equation

- a. $pH = -\log [H^+]$ (ppm 1.4.1 and 1.4.2)
- b. $[H^+] = 10^{-pH}$ (ppm 1.4.1 and 1.4.2)

Part A Completion

Write the answers to the following questions in the spaces provided. Show necessary calculations and units for calculations involving either mass, volume, or moles.

1. Acetic acid, CH_3COOH , is a weak acid. K_a for acetic acid is 1.8×10^{-5} . What is the pH of a 0.10 M solution of acetic acid? a. 4.74
2. What is the pH of a 0.10 M solution of sodium acetate, CH_3COONa ? b. 9.26
3. What is the pH of a 0.10 M solution of hydrochloric acid, HCl ? c. 1.00
4. What is the pH of a 0.10 M solution of sodium hydroxide, $NaOH$? d. 13.00
5. What is the pH of a 0.10 M solution of acetic acid, CH_3COOH ? e. 2.87
6. What is the pH of a 0.10 M solution of sodium acetate, CH_3COONa ? f. 8.13
7. What is the pH of a 0.10 M solution of hydrochloric acid, HCl ? g. 1.00
8. What is the pH of a 0.10 M solution of sodium hydroxide, $NaOH$? h. 13.00
9. What is the pH of a 0.10 M solution of acetic acid, CH_3COOH ? i. 2.87
10. What is the pH of a 0.10 M solution of sodium acetate, CH_3COONa ? j. 8.13

Hydrogen Ions And Acidity Section Review Answers

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