

Mark Scheme

Q1.

Question Number	Acceptable Answer	Additional Guidance	Mark
(a) (i)	$\text{CaCO}_3(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$ Balanced equation (1) State symbols (1)	Accept $\text{CaCO}_3(\text{s}) + 2\text{H}^+(\text{aq}) \rightarrow \text{Ca}^{2+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$ 2nd mark dependent on first or near miss. Reject $\text{H}_2\text{CO}_3(\text{aq})$ in equation, but allow state symbol mark if otherwise correct.	(2)
Question Number	Acceptable Answer	Additional Guidance	Mark
(a) (ii)	Finds molar mass of calcium carbonate	Example of calculation $\text{Mr of calcium carbonate}$ $= 40.1 + 12 + (16 \times 3) = 100.1 \text{ (g mol}^{-1}\text{)}$ Allow $= 40 + 12 + (16 \times 3) = 100 \text{ (g mol}^{-1}\text{)}$ Accept answer with no working	(1)

May 2012 Chemistry Mark Scheme Edexcel

Rachel Sandford



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