

Sulfur clock lab

Experimental procedure

1. Measurement of the formation of visible product
2. Measurement of the initial concentration of reactants
3. Prepare solutions A, B, C, D and E as shown:

see Table 10.10.10.10

	Reactants	Reagents	Apparatus
A	1 mL	0.1 M H_2O_2 , 0.01 M KSCN	Reaction setup
B	0.5 mL	0.1 M H_2O_2 , 0.01 M KSCN	0.5 mL, 0.01 M H_2O_2 , 0.01 M KSCN
C	0.25 mL	0.1 M H_2O_2 , 0.01 M KSCN	0.25 mL, 0.01 M H_2O_2 , 0.01 M KSCN
D	0.125 mL	0.1 M H_2O_2 , 0.01 M KSCN	0.125 mL, 0.01 M H_2O_2 , 0.01 M KSCN

4. Add the solution using a graduated cylinder with completely standard chemical formulae for the reaction.
5. Transfer exactly the calculated amount of H_2O_2 solution to the reaction mixture using a graduated cylinder.
6. Add exactly 0.01 M KSCN solution to the reaction mixture.
7. Prepare and use the standard solutions.
8. Prepare solutions A, B, C, D and E as shown:
 - a. Transfer 0.1 mL of 0.1 M H_2O_2 solution to the reaction mixture.
 - b. Add 0.01 M KSCN solution to the reaction mixture using a graduated cylinder.
 - c. Dilute to the 10 mL mark using distilled water.
 - d. Use the flask and record the initial concentration.
9. Transfer solution A into the graduated cylinder as 0.1 mL and transfer the rest of the solution into any other flask.

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