

The Hydrogen Atom – Student Challenge

Background Material

Carefully read the background-information Energy Levels, Light, and Transitions, and answer the following questions as clearly (and completely) as you can.

Operation 1: Compare the following table which compares how the Bohr Model and the Quantum-mechanical approach the Hydrogen atom. In each “describing” box describe things as accurately as you can assuming they are true.

Concept	Bohr Model	Quantum Model
Electrons	positively charged particles in flat orbits	positively charged particles in flat orbits
Electrons	orbiting electrons	negatively charged clouds
Orbitals	Electrons do not move, they give the probability picture	Electrons, being particles without mass
Transitions of electrons to the first orbital	using a formula to find the radius	Classical wave
Energy of the lowest energy level	$E_1 = -13.6 \text{ eV}$	It is about?

Operation 2: Compare the following table which compares the properties of two different elements by using their approximate relationships. The data now is completed for you. The red element has a higher frequency, smaller frequency, smaller energy, and the same wavelength. Which space is which element?

Hydrogen Atom Student Guide Solutions Naap

Puja Mehta

A decorative red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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