

Acids and bases

Chapter outline

- Recall the definitions of Brønsted–Lowry acids and bases and deduce these in a chemical reaction, including the identification of conjugate acid–base pairs.
- Recall the terms ‘amphoteric’ and ‘amphiprotic’.
- Describe the reactions of acids with metals, metal oxides, metal hydroxides, hydrogen carbonates and carbonates.
- Recall that pH is a measure of the concentration of aqueous hydrogen ions and solve problems involving pH and the concentration of hydrogen ions and hydroxide ions.
- Explain the difference between strong and weak acids and bases and describe experiments that could be used to distinguish these.
- Recall what is meant by acid deposition and explain some of the problems associated with acid deposition as well as pre- and post-combustion methods of reducing sulfur oxide emissions.
- Recall the definitions of Lewis acids and bases and recognise these in a chemical reaction. **HL**
- Deduce the dissociation constant for weak acids and bases. **HL**
- Sketch and explain the important features of titration curves including the composition and action of a buffer and use these curves to deduce the pK_a or pK_b value of a weak acid or base. **HL**
- Explain how an acid–base indicator works and select a suitable indicator for a titration. **HL**

KEY TERMS

Acid deposition: A more general term than acid rain; it refers to any process in which acidic substances leave the atmosphere and are deposited on the Earth’s surface.

Brønsted–Lowry acid: A proton (H^+) donor.

Brønsted–Lowry base: A proton (H^+) acceptor.

pH: A measure of the concentration of hydrogen ions in an aqueous solution.
 $pH = -\log_{10}[H^+]$

Strong acid/base: An acid/base that is completely dissociated into its ions in aqueous solution.

Weak acid/base: An acid/base that is only partially dissociated into its ions in aqueous solution.

Guided Reading And Study Workbook Chapter 20 Acids Bases

John L. Cameron, Andrew M. Cameron



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Table of Contents Guided Reading And Study Workbook Chapter 20 Acids Bases

1. Understanding the eBook Guided Reading And Study Workbook Chapter 20 Acids Bases
 - The Rise of Digital Reading Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Advantages of eBooks Over Traditional Books
2. Identifying Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guided Reading And Study Workbook Chapter 20 Acids Bases
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Personalized Recommendations
 - Guided Reading And Study Workbook Chapter 20 Acids Bases User Reviews and Ratings
 - Guided Reading And Study Workbook Chapter 20 Acids Bases and Bestseller Lists

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 - Adjustable Fonts and Text Sizes of Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Highlighting and Note-Taking Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Interactive Elements Guided Reading And Study Workbook Chapter 20 Acids Bases
8. Staying Engaged with Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guided Reading And Study Workbook Chapter 20 Acids Bases
9. Balancing eBooks and Physical Books Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guided Reading And Study Workbook Chapter 20 Acids Bases
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Setting Reading Goals Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Fact-Checking eBook Content of Guided Reading And Study Workbook Chapter 20 Acids Bases
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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