

## Honors Final Review

*Note: this is in no way guaranteed to be comprehensive. Anything from the year is fair game, but the emphasis will be on more recent material. Bolded items are most important.*

### Chapter 1: Introductory stuff.

- Qualitative vs quantitative measurements.
- Scientific method

### Chapter 2: Atoms, elements, and mixtures.

- Matter is made of up atoms
- **Elements, compounds, and molecules: elements are only one type of atom, compounds have multiple elements, and molecules are atoms that are stuck together.**
- States of matter—solid, liquid, and gas—and their properties.
- Physical versus chemical changes.
- Mixtures: homogeneous and heterogeneous.

### Chapter 3: The elements and ions.

- Dalton's atomic theory: elements are made of atoms, all atoms of a given element are identical, atoms of different elements are different, and reactions don't alter the atoms—only rearrange them.
- Structure of the atom: protons and neutrons in the nucleus, electrons around the outside.
- J.J. Thomson discovered electrons.
- Rutherford shot particles at gold foil and saw some scatter, some go through, meaning that part of the atom was hard while most was empty space.
- Element symbols: atomic number and atomic mass
- **Atomic number = number of protons, defines an element.**
- **Atomic mass = # protons + # neutrons**
- **Isotopes have same number of protons, different number of neutrons (same element, but weigh different amounts).**
- **Number of electrons changes, leading to charges. Things with a charge are called ions.**
- Positively charged atoms/groups of atoms are cations. Negatively charged ones are anions.
- Periodic table: alkali metals, alkaline earth metals, transition metals, metalloids, nonmetals, halogens, and noble gases.
- **Certain groups of elements always form ions with the same charge. See page 74.**
- Compounds containing ions conduct electricity if the ions can move.
- **Writing formulas for ionic compounds: the charges have to balance out to neutral. Use the crossing method with the charges, but reduce if there's a common factor. For example:**  
 $\text{Mg}^{2+}$  and  $\text{N}^{3-}$  for  $\text{Mg}_3\text{N}_2$       but       $\text{Mg}^{2+}$  and  $\text{O}^{2-}$  form  $\text{MgO}$

### Chapter 4: Naming compounds.

- **Ionic compounds have a metal and at least one non-metal, and are stuck together by charges. Covalent compounds have only non-metals, and are held together by bonds.**
- Anions of a single element use the start of the element name, and end in *-ide*. Nitrogen with a negative charge becomes nitride. Oxygen with a negative charge is oxide.
- Ions can be a group of elements all stuck together by covalent bonds, that overall have a charge.
- If the group has an element attached to some number of oxygens, then it ends in *-ate*, and starts with the beginning of the name of the non-oxygen element. For example,  $\text{NO}_3^-$  is nitrate,  $\text{PO}_4^{3-}$  is phosphate.
- You can have a different number of oxygens, which have different naming schemes. The one that ends in *-ate* is the most common one.

# Mythology And Legends Final Exam Study Guide

**Camilla Rothe**



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