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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal kbr that contains 2 92 kg of water 1 molality moles solute divided by kilograms solute 0 449 mol kg x 2 92 kg x 1 31108 mol of kbr 2 moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g kbr

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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