

19.5

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the requested unit or numerical information of the values provided that are indicated in the correct space. Write numbers in scientific notation when appropriate.

1. What substance is _____, a strong hydrogen ion donor?
2. _____ is the concentration of H^+ ions present in 0.001 M HCl solution.
3. _____ is the concentration of H^+ ions in 0.001 M HCl solution.
4. _____ is the concentration of H^+ ions in 0.001 M HCl solution.
5. _____ is the concentration of H^+ ions in 0.001 M HCl solution.
6. _____ is the concentration of H^+ ions in 0.001 M HCl solution.
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12. _____ is the concentration of H^+ ions in 0.001 M HCl solution.
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15. _____ is the concentration of H^+ ions in 0.001 M HCl solution.
16. _____ is the concentration of H^+ ions in 0.001 M HCl solution.
17. _____ is the concentration of H^+ ions in 0.001 M HCl solution.
18. _____ is the concentration of H^+ ions in 0.001 M HCl solution.
19. _____ is the concentration of H^+ ions in 0.001 M HCl solution.
20. _____ is the concentration of H^+ ions in 0.001 M HCl solution.

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Hydrogen Ions And Acidity Section Review Answers

DP Hallahan



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