

## Virtual Lab: Empirical Formula of a Hydrate

### Background:

Hydrates are ionic compounds (salts) that have a definite amount of water (water of hydration) as part of their structure. The water is chemically combined with the salt in a **definite ratio**. Ratios vary in different hydrates but are specific for any given hydrate.

The formula of a hydrate is represented in a special manner. The hydrate of copper sulfate in this experiment is listed below:



In the formula, the unit formula for the salt appears first, and the water formula is last. *The raised dot means that the water is loosely bonded to the salt.* The coefficient  $x$  stands for the number of molecules of water bonded to one unit of salt. This special formula, like all other formulas, illustrates the law of definite composition.

When hydrates are heated, the "water of hydration" is released as vapor. The remaining solid is known as the **anhydrous salt**. The general reaction for heating a hydrate is:



Some anhydrous salts can absorb moisture from the air to become hydrated. A common example is silica gel, made from sodium silicate which is usually packaged with optical or electronic devices shipped by boat. The silica gel protects the devices from high humidity.



We can find the percent of water in a hydrate experimentally by accurately determining the mass of the hydrate and the mass of the anhydrous salt. The difference in mass is due to the water lost by the hydrate. The percentage of water in the original hydrate can easily be calculated:

$$\text{Percent H}_2\text{O} = \frac{\text{mass of H}_2\text{O}}{\text{mass of hydrate}} \times 100$$

# Hydrates Lab Answers

**Joshua Romine Morton, Donald Roop  
Clippinger, Lawrence Powell Eblin**

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Unilever Research Laboratory Colworth Welwyn Colworth House Sharnbrook Bedford January 1973 Contents Contents of Volume 1 xv Contents of Volume 3 xvi Contents of Volume 4 xvii Chapter 1 The Solvent Properties of Water F Franks 1 Water the Universal Solvent the Study of Aqueous Solutions 2 Aqueous Solutions of Nonelectrolytes 5 2 1 Apolar Solutes 6 2 2 Polar Solutes 19 2 3 Ionic Solutes Containing Alkyl Residues Apolar Electrolytes 38 3 Aqueous Solutions of Electrolytes 42 3 1 Single Ion Properties 42 3 2 Ion Water Interactions 43 3 3 Interionic Effects 47 4 Complex Aqueous Mixtures 48 Chapter 2 Water in Stoichiometric Hydrates M Falk and O Knop 1 Introduction 55 2 Symmetry and Types of Environment of the H<sub>2</sub>O Molecule 2 in Crystals 57 vii Contents viii 2 1 Site Symmetry 57

*Unconventional Fuels Part 2: The Promise of Methane Hydrates, Serial No. 111-32, July 30, 2009, 111-1 Oversight Hearing, \**,2010

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faculty of higher education researchers and academicians      **S. 1418, the Methane Hydrate Research and Development Act of 1998** United States. Congress. House. Committee on Science. Subcommittee on Energy and Environment,1999      *Nuclear Science Abstracts* ,1972      *Instructors Manual to Lab Manual* Ralph Petrucci,William Harwood,Geoffrey Herring,2001      Rock Products ,1926      **Subject Index to Unclassified ASTIA Documents** Defense Documentation Center (U.S.),1960      **My One-Eyed, Three-Legged Therapist** Kathy M. Finley,2023-10-15 My One Eyed Three Legged Therapist How My Cat Clio Saved Me is the story of how an adorable spunky gray and white kitten helped the author regain the courage to face life s challenges and realize that none of us is truly alone Born into poverty losing her dad at age seven and targeted by bullies Kathy turned to pets for unconditional love and acceptance A difficult childhood led to an abusive marriage but things changed on her fortieth birthday when her staff at the organization where she worked gave her an extraordinary cat named Clio The runt of the litter a two time cancer survivor and a special needs cat Clio nevertheless had an incredible will to live full tilt This intrepid feline knew no fear and displayed unlimited self confidence She overcame not one but two disabilities By watching Clio thrive despite what life threw at her Kathy was able to put her own life in perspective by learning to accept the past embrace the present and look forward to the future      *Energy Research Abstracts* ,1992

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