

## Notes on Study-Session

### Changes in Kinetics:

Rate of Reaction:

Rate of Consumption:  $\Delta[\text{Reactant}]$

Rate of Change:  $\Delta[\text{Product}]$

Rate of Consumption:  $\text{Average Reaction Rate} = \frac{\Delta[\text{Reactant}]}{\Delta t}$  and  $\Delta[\text{Product}]$  Experimentally (Reaction Time)

### Rate Information will be:

- instantaneous phase diagram
- position of initial conditions
- position of final conditions
- distinguish between initial and final conditions
- instantaneous reaction rate
- initial & final conditions of reaction

### Rate of Reaction:

(1)  $\text{Rate of Reaction} = \frac{\Delta[\text{Reactant}]}{\Delta t}$

**Reaction Rate:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

**Initial:** - the concentration of reactants or products at the start of the reaction

**Final:** - the concentration of reactants or products at the end of the reaction

**Rate:** - the change in concentration over time

**Instantaneous:** - the concentration of reactants or products at a specific time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

**Position:** - the concentration of reactants or products at a specific time

**Initial rate:** - the concentration of reactants or products at the start of the reaction

**Final rate:** - the concentration of reactants or products at the end of the reaction (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

**Reaction rate:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

(2) **Rate of Reaction:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

(3) **Rate of Reaction:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

(4) **Rate of Reaction:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

(5) **Rate of Reaction:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )



(6) **Rate of Reaction:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

(7) **Rate of Reaction:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

(8) **Rate of Reaction:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

(9) **Rate of Reaction:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

(10) **Rate of Reaction:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

(11) **Rate of Reaction:** - the change in concentration of reactants or products over time (usually expressed in  $\text{mol dm}^{-3} \text{s}^{-1}$ )

# Myers Unit 10 Study Guide Answers

**S Marginson**



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