

A student investigates the evaporation of water. He pours  $100\text{cm}^3$  of water into measuring cylinder B. Measuring cylinder A is kept at  $40^\circ\text{C}$  and B is kept at  $80^\circ\text{C}$  in the same part of the laboratory. Fig. 2.1 shows the two measuring cylinders after 3 days.



**Fig. 2.1**

- (a) (i) Using ideas about molecules, explain why more water evaporates when the water is kept at  $80^\circ\text{C}$ .

.....  
 .....  
 .....  
 .....  
 .....  
 ..... [2]

- (ii) Apart from an increase in temperature, state one change that causes water to evaporate faster.

.....  
 ..... [1]

- (b) The specific latent heat of vaporisation of water is  $2300\text{J/g}$  and the density of water is  $1.0\text{g/cm}^3$ . During the three days, the water level in B drops from the  $100\text{cm}^3$  mark to the level shown in Fig. 2.1. Calculate the energy used to evaporate water from B during the three days.

energy = ..... [2]

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**Richard Bailey**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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