

Summative Test in Grade 10 Science

Name: _____

Score: _____

Name of Teacher: _____

Section: _____

Instruction: Write the letter of the correct answer before the number.

1. Chemical reactions _____

- A. Occur only in living things
B. Produce new substances
C. create and destroy atoms
D. only occur outside living organisms

2. What are the coefficients that will balance the skeleton equation? $\text{AlCl}_3 + \text{NaOH} \rightarrow \text{Al(OH)}_3 +$

NaCl

- A. 1,3,1,3
B. 3,1,3,1
C. 1,1,1,3
D. 1,3,3,1

3. In every balanced chemical equation, each side of the equation has the same number of _____

- A. Atoms in each element
B. molecules
C. moles
D. coefficients

4. Which molecules are most closely related to lipids?

- A. nucleotides
B. fatty acids
C. amino acids
D. sugars

5. Which of the following molecules are classified as carbohydrates?

- A. nucleotides
B. fatty acids
C. amino acids
D. sugars

6. Polysaccharides are _____

- A. lipids
B. carbohydrates
C. proteins
D. nucleic acids

7. Which molecules are most closely related to proteins?

- A. nucleotides
B. fatty acids
C. amino acids
D. sugars

8. Which of the following is not an example of matter?

- A. Air
B. heat
C. smoke
D. water vapor

9. When the equation $\text{Fe} + \text{Cl}_2 \rightarrow \text{FeCl}_3$ is balanced, what is the coefficient for Cl_2 ?

- A. 1
B. 2
C. 3
D. 4

10. What are the missing coefficients for the skeleton equation below?



- A. 1,3,2,3
B. 2,12,4,6
C. 4,6,2,3
D. 1,6,2,3

11. The equation above is an example of what type of chemical reaction?

- A. combination
B. decomposition
C. single displacement
D. double displacement

12. The type of reaction that takes place when one element reacts with a compound to form a new compound and a different element is a _____ reaction.

- A. combination
B. decomposition
C. single displacement
D. double displacement

13. Which of the following is a balanced equation representing the decomposition of lead(IV) oxide?

- A. $\text{PbO}_2 \rightarrow \text{Pb} + 2\text{O}$
B. $\text{PbO}_2 \rightarrow \text{Pb} + \text{O}_2$
C. $\text{Pb}_2\text{O} \rightarrow 2\text{Pb} + \text{O}$
D. $\text{PbO} \rightarrow \text{Pb} + \text{O}_2$

14. Boyle's law states that _____

- A. the volume of a gas varies inversely with pressure
B. the volume of a gas varies directly with pressure
C. the temperature of a gas varies inversely with pressure
D. the temperature of a gas varies directly with pressure

15. Charles's law states that _____

- A. the pressure of a gas is inversely proportional to its temperature in kelvins
B. the volume of a gas is directly proportional to its temperature in kelvins
C. the pressure of a gas is directly proportional to its temperature in kelvins
D. the volume of a gas is inversely proportional to its temperature in kelvins

16. As the temperature of a fixed volume of a gas increases, the pressure will _____

- A. Vary inversely
B. decrease
C. not change
D. increase

17. A chemical reaction is a process in which _____

- A. products change into reactants
B. the law of conservation of mass applies
C. substances change state
D. all of these

18. During a chemical reaction,

- A. new elements are produced
B. atoms are rearranged
C. atoms are destroyed
D. elements are destroyed

19. An equation is balanced by,

- A. changing subscripts
B. adding coefficients
C. erasing elements as necessary
D. adding elements as necessary

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