Guizizz	NAME:	
Stoichiometry Practice 10 Questions	CLASS:	
	DATE :	

1	Balance the following reaction:						
	Mg(s) + HCl(aq)> MgC	:I ₂ (aq) +	$H_2(g)$				

- 1.2.1.2 1,2,2,2
- 2,1,1,2 1,2,1,1
- 2 Using the balanced equation in question 1, $Mg(s) + HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$ How many grams of HCI are consumed by the reaction of 2.50 moles of magnesium?
- A. 218 g 7.29
- 5.00 g 182 g
 - Using the balanced equation in question 1, $Mg(s) + HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$ What is the mass in grams of H₂ gas when 4.0 moles of HCI is added to the reaction?
- A $0.99\,\mathrm{g}$ 4.0 g C 2.0 g 6.0 g
- Using the balanced equation in question 1, $Mg(s) + HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$ How many moles of HCI are consumed in the production of 7.5 moles of MgCl₂?

7.5 moles:

- 15 moles 3.8 moles C (D)
- Balance the following reaction: $CaC_2(s) + H_2O(l) --> C_2H_2(g) + Ca(OH)_2(aq)$

23 moles

- 2, 1, 1, 1 2,1,2,1
- 1,2,2,2 1,2,1,1

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- 1. Understanding the eBook Grade 11 Physics Exam Questions Stoichiometry
 - The Rise of Digital Reading Grade 11 Physics Exam Questions Stoichiometry
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Grade 11 Physics Exam Questions Stoichiometry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 11 Physics Exam Questions Stoichiometry
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Grade 11 Physics Exam Questions Stoichiometry

- Personalized Recommendations
- o Grade 11 Physics Exam Questions Stoichiometry User Reviews and Ratings
- Grade 11 Physics Exam Questions Stoichiometry and Bestseller Lists
- 5. Accessing Grade 11 Physics Exam Questions Stoichiometry Free and Paid eBooks
 - Grade 11 Physics Exam Questions Stoichiometry Public Domain eBooks
 - Grade 11 Physics Exam Questions Stoichiometry eBook Subscription Services
 - o Grade 11 Physics Exam Questions Stoichiometry Budget-Friendly Options
- 6. Navigating Grade 11 Physics Exam Questions Stoichiometry eBook Formats
 - o ePub, PDF, MOBI, and More
 - Grade 11 Physics Exam Questions Stoichiometry Compatibility with Devices
 - Grade 11 Physics Exam Questions Stoichiometry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 11 Physics Exam Questions Stoichiometry
 - Highlighting and Note-Taking Grade 11 Physics Exam Questions Stoichiometry
 - Interactive Elements Grade 11 Physics Exam Questions Stoichiometry
- 8. Staying Engaged with Grade 11 Physics Exam Questions Stoichiometry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 11 Physics Exam Questions Stoichiometry
- 9. Balancing eBooks and Physical Books Grade 11 Physics Exam Questions Stoichiometry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 11 Physics Exam Questions Stoichiometry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Grade 11 Physics Exam Questions Stoichiometry
 - Setting Reading Goals Grade 11 Physics Exam Questions Stoichiometry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade 11 Physics Exam Questions Stoichiometry

- Fact-Checking eBook Content of Grade 11 Physics Exam Questions Stoichiometry
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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