# Mini-Lab

# Mole and Stoichiometry Mini-Lab

 Write and balance the equation for the reaction between aluminum and water to produce aluminum oxide and hydrogen gas.

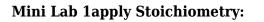
2 AAIL++ 37HGO-++ AAIGO3 ++ 7B2 H2

- 8. Using the moles of the aluminum can in #1 as the GIVEN, how many moles of hydrogen gas can be produced? Assume water is available in excess. <u>UNKNOWN</u>: mol H<sub>2</sub> GIVEN: mol Al (answer to #1)
- Using the moles of water from #4 as the GIVEN with aluminum in excess, calculate the number of moles of aluminum oxide that can be produced.

UNKNOWN: molALO3 GIVEN: mol H2O (answer to #4)

# **Mini Lab 1apply Stoichiometry**

**Irene Cesa** 



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