

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.

$$\Delta S \geq 0$$

 where ΔS is the change in entropy.



- **Enthalpy** (H) is a thermodynamic property that represents the total energy of a system, including internal energy and the energy required to displace the environment.

$$H = U + PV$$

 where U is internal energy, P is pressure, and V is volume.
 - **Enthalpy Change** (ΔH) is the change in enthalpy during a process.

$$\Delta H = H_{\text{final}} - H_{\text{initial}}$$

 - **Heat Capacity** (C) is the amount of heat required to raise the temperature of a substance by a certain amount.

$$Q = C \Delta T$$

 where Q is heat, C is heat capacity, and ΔT is the change in temperature.



- **Free Energy** (G) is a thermodynamic property that represents the maximum work that can be extracted from a system at constant temperature and pressure.

$$G = H - TS$$

 where H is enthalpy, T is temperature, and S is entropy.
 - **Free Energy Change** (ΔG) is the change in free energy during a process.

$$\Delta G = \Delta H - T \Delta S$$

 where ΔH is the change in enthalpy, T is the temperature, and ΔS is the change in entropy.

- **Heat Engines** are devices that convert heat into work.
 - **Efficiency** (η) is the ratio of the work output to the heat input.

$$\eta = \frac{W}{Q_{\text{in}}}$$

 where W is work and Q_{in} is the heat input.



- **Thermodynamic Processes** are changes in the state of a system.
 - **Isobaric Process**: A process that occurs at constant pressure.

$$P = \text{constant}$$

 - **Isochoric Process**: A process that occurs at constant volume.

$$V = \text{constant}$$

 - **Isothermal Process**: A process that occurs at constant temperature.

$$T = \text{constant}$$

 - **Adiabatic Process**: A process that occurs without heat exchange with the surroundings.

$$Q = 0$$

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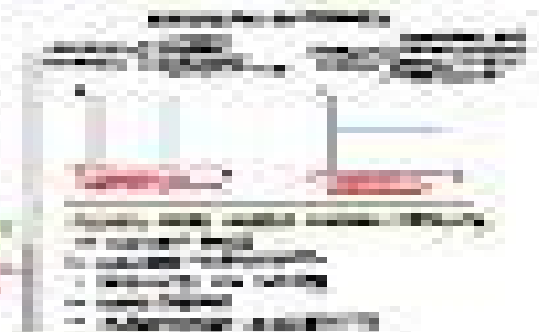
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Microeconomics Final Study Guide

David Kirk



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