



K5 Blazer Manual Transmission Conversion

Peipei Pang



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 what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is
 a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 $\text{Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$
 oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ 24 3 $\text{Ag} + \text{Ni}^{3+} \rightarrow \text{Ni} + 3\text{Ag}$ oxidation reduction 25 $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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 magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a $\text{I}_2 + \text{O}_5$
 $5\text{Co} + \text{I}_2 \rightarrow 5\text{Co}^{2+} + 2\text{I}^-$ 5 c o i 2 5 c o 2

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 reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation
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 reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ $\text{Sr} \rightarrow \text{Sr}^{2+}$ oxidized reducing agent $\text{O}_2 \rightarrow \text{O}^{2-}$ reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ $\text{Li} \rightarrow \text{Li}^+$ oxidized red ag
 $\text{S} \rightarrow \text{S}^{2-}$ reduced ox ag $3\text{Cs} + \text{Br}_2$

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