

Matter, Chemical Trends and Chemical Bonding Unit Test

Knowledge	Thinking	Application	Communication
11/14	12/17	13/16	12/12

Thinking/Investigation (17 marks)

1. An element is composed of 70% of an isotope with an atomic mass of 69.0u. If the other isotope has 2 more neutrons per atom, what is the average atomic mass of the element (4 marks)?

$$AAM = (M_1 \times \text{Abundance}) + (M_2 \times \text{Abundance})$$

$$AAM = (69.0u \times 70\%) + (71.0u \times 30\%)$$

$$AAM = 69.6u$$

4

2. Fill in the following chart (13 marks)

Formula	Lewis Structure	Name(s) of Shape	Overall Polarity of Molecule
SBr_2	$\text{:}\ddot{\text{Br}}\text{--}\ddot{\text{S}}\text{--}\ddot{\text{Br}}\text{:}$ $\text{:}\ddot{\text{Br}}\text{:}$	Trigonal planar ✓	Non-Polar ✓
PO_3^{3-}	$\left[\text{:}\ddot{\text{O}}\text{--}\ddot{\text{P}}\text{--}\ddot{\text{O}}\text{:} \right]^{-3}$ $\text{:}\ddot{\text{O}}\text{:}$	Trigonal pyramidal ✓	Non-Polar ✗
HCP	$\text{H--}\ddot{\text{C}}\equiv\text{P:}$ why so many ✓	Bent ✗	Non-Polar ✓
CH_3NH_2	$\begin{array}{c} \text{H} & \text{H} \\ & \\ \text{H--C--N:} \\ & \\ \text{H} & \text{H} \end{array}$ ✓	Trigonal planar ✓	Non-Polar ✗

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