EMPIRICAL AND MOLECULAR FORMULAE

EMRIRICAL FORMULA - This represents the simplest whole number ratio of atoms within a molecule

MOLECULAR FORMULA - This represents the exact numbers of each atom in a molecule

Empirical formulae can be calculated from either % mass composition data or mass composition data. Molecular formula can then be calculated if the molecular mass of the compound is known.

EXAMPLE 1 - from % composition

A substance containing carbon and hydrogen and oxygen only contains 73.47% carbon and 10.20% hydrogen by mass, its relative molecular mass is 98. Calculate its molecular formula.

Step 1 calculate empirical formula

ATOMS	CARBON	HYDROGEN	OXYGEN
PERCENTAGE MASS	73.47	10.20	(100 - (73.47 + 10.20) = 16.33
ATOMIC MASS	12	1	16
MASS/ATOMIC MASS	6.1225	10.20	1.020
RATIO(divide by smallest)	6	10	1

Empirical formula = C.H.O

Step 2 use Mo to find molecular formula

 $M_{\rm P}$ = mass ($C_6H_{10}O$) \times n

98 = (98) x n n = 1

molecular formula =

CHO

EXAMPLE 2 - from mass data

A substance contains 12.28g nitrogen, 3.51g hydrogen, 28.07g sulphur and 56.10g oxygen, its Me is 228 .Calculate its molecular formula.

Empirical And Molecular Answers And Questions

B Lingard

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