

## Mark Scheme

Q1.

Question Number	Acceptable Answer	Additional Guidance	Mark
(a) (i)	$\text{CaCO}_3(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$ <p>Balanced equation (1)</p> <p>State symbols (1)</p>	<p>Accept  <math display="block">\text{CaCO}_3(\text{s}) + 2\text{H}^+(\text{aq}) \rightarrow \text{Ca}^{2+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})</math></p> <p>2nd mark dependent on first or near miss.</p> <p>Reject <math>\text{H}_2\text{CO}_3(\text{aq})</math> in equation, but allow state symbol mark if otherwise correct.</p>	(2)
Question Number	Acceptable Answer	Additional Guidance	Mark
(a) (ii)	Finds molar mass of calcium carbonate	<p>Example of calculation  <math display="block">\text{Mr of calcium carbonate} = 40.1 + 12 + (16 \times 3) = 100.1 \text{ (g mol}^{-1}\text{)}</math>                     Allow  <math display="block">= 40 + 12 + (16 \times 3) = 100 \text{ (g mol}^{-1}\text{)}</math>                     Accept answer with no working</p>	(1)

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**S Marginson**



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