

# 7 Ionic and Metallic Bonding

## Big idea BONDING AND INTERACTIONS

### 7.1 Ions

For students using the Foundation edition, assign problems 1, 3–5, 7–12, 14, 15, 18–20.

**Essential Understanding** Ions form when atoms gain or lose valence electrons, becoming electrically charged.

#### Lesson Summary

**Valence Electrons** Valence electrons are the electrons in the outermost occupied energy level and are involved in ion formation.

- ▶ For a representative element, the group number equals the number of valence electrons the atom contains.
- ▶ An electron dot structure shows the symbol of the element and its valence electrons.
- ▶ Atoms tend to gain or lose the number of electrons that will provide the atom with a noble gas electron configuration.

**Formation of Cations** Cations are positively charged ions formed when an atom loses one or more valence electrons.

- ▶ Atoms and the cations formed from them have different properties.
- ▶ Elements in Group 1 form cations with a charge of 1+, and those in Group 2 form cations with a charge of 2+.
- ▶ Many transition metals form more than one cation and do not follow the octet rule.

**Formation of Anions** Anions are negatively charged ions formed when an atom gains one or more valence electrons.

- ▶ Commonly, the name of an anion ends in *-ide*.
- ▶ Anions form from nonmetallic elements.
- ▶ The anions formed from halogens are known as halides.

After reading Lesson 7.1, answer the following questions.

#### Valence Electrons

1. What are valence electrons?

*Valence electrons are the electrons in the highest occupied energy level of an element's atoms.*

2. The valence electrons largely determine the *chemical properties* of an element and are usually the only electrons used in *chemical bonds*.

# Lesson 5 1 Reading And Study Workbook

**Rachel S Tattersall**



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## **Table of Contents Lesson 5 1 Reading And Study Workbook**

1. Understanding the eBook Lesson 5 1 Reading And Study Workbook
  - The Rise of Digital Reading Lesson 5 1 Reading And Study Workbook
  - Advantages of eBooks Over Traditional Books
2. Identifying Lesson 5 1 Reading And Study Workbook
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Lesson 5 1 Reading And Study Workbook
  - User-Friendly Interface
4. Exploring eBook Recommendations from Lesson 5 1 Reading And Study Workbook

- Personalized Recommendations
  - Lesson 5 1 Reading And Study Workbook User Reviews and Ratings
  - Lesson 5 1 Reading And Study Workbook and Bestseller Lists
5. Accessing Lesson 5 1 Reading And Study Workbook Free and Paid eBooks
- Lesson 5 1 Reading And Study Workbook Public Domain eBooks
  - Lesson 5 1 Reading And Study Workbook eBook Subscription Services
  - Lesson 5 1 Reading And Study Workbook Budget-Friendly Options
6. Navigating Lesson 5 1 Reading And Study Workbook eBook Formats
- ePub, PDF, MOBI, and More
  - Lesson 5 1 Reading And Study Workbook Compatibility with Devices
  - Lesson 5 1 Reading And Study Workbook Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Lesson 5 1 Reading And Study Workbook
  - Highlighting and Note-Taking Lesson 5 1 Reading And Study Workbook
  - Interactive Elements Lesson 5 1 Reading And Study Workbook
8. Staying Engaged with Lesson 5 1 Reading And Study Workbook
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Lesson 5 1 Reading And Study Workbook
9. Balancing eBooks and Physical Books Lesson 5 1 Reading And Study Workbook
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Lesson 5 1 Reading And Study Workbook
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Lesson 5 1 Reading And Study Workbook
- Setting Reading Goals Lesson 5 1 Reading And Study Workbook
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lesson 5 1 Reading And Study Workbook

- Fact-Checking eBook Content of Lesson 5 1 Reading And Study Workbook
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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