

## Naming Compounds Containing Polyatomic Ions

Use the following list of polyatomic ions to name the compounds given.

OH  
 NO<sub>3</sub>  
 C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>

hydroxide  
 nitrate  
 acetate

ClO<sub>3</sub>  
 HCO<sub>3</sub>  
 SO<sub>4</sub>  
 CO<sub>3</sub>

chlorate  
 bicarbonate  
 sulfate  
 carbonate

SO<sub>3</sub>  
 PO<sub>4</sub>  
 NH<sub>4</sub>

sulfite  
 phosphate  
 ammonium

- |  |   |
|--|---|
| 1. CuSO <sub>4</sub> _____                             | 26. NaOH _____  |
| 2. KClO <sub>3</sub> _____                             | 27. NaNO <sub>3</sub> _____                               |
| 3. NH <sub>4</sub> OH _____                            | 28. Ba(OH) <sub>2</sub> _____                             |
| 4. K <sub>2</sub> CO <sub>3</sub> _____                | 29. K <sub>2</sub> CO <sub>3</sub> _____                  |
| 5. Na <sub>2</sub> SO <sub>4</sub> _____               | 30. LiOH _____  |
| 6. KC <sub>2</sub> H <sub>3</sub> O <sub>2</sub> _____ | 31. KHCO <sub>3</sub> _____                               |
| 7. NH <sub>4</sub> Cl _____                            | 32. NH <sub>4</sub> F _____                               |
| 8. Mg(OH) <sub>2</sub> _____                           | 33. KClO <sub>3</sub> _____                               |
| 9. Zn(NO <sub>3</sub> ) <sub>2</sub> _____             | 34. NaC <sub>2</sub> H <sub>3</sub> O <sub>2</sub> _____  |
| 10. MgSO <sub>4</sub> _____                            | 35. CaClO <sub>3</sub> _____                              |
| 11. CaCO <sub>3</sub> _____                            | 36. SrNO <sub>3</sub> _____                               |
| 12. KOH _____  | 37. Be(OH) <sub>2</sub> _____                             |
| 13. Ca(OH) <sub>2</sub> _____                          | 38. Li <sub>2</sub> SO <sub>4</sub> _____                 |
| 14. K <sub>2</sub> CO <sub>3</sub> _____               | 39. Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> _____ |
| 15. AlPO <sub>4</sub> _____                            | 40. BaSO <sub>3</sub> _____                               |
| 16. Na <sub>2</sub> PO <sub>4</sub> _____              | 41. NH <sub>4</sub> I _____                               |
| 17. MgSO <sub>4</sub> _____                            | 42. NH <sub>4</sub> HCO <sub>3</sub> _____                |
| 18. AgNO <sub>3</sub> _____                            | 43. Cu <sub>2</sub> (PO <sub>4</sub> ) <sub>3</sub> _____ |
| 19. Na <sub>2</sub> CO <sub>3</sub> _____              | 44. Ni(OH) <sub>2</sub> _____                             |
| 20. NaHCO <sub>3</sub> _____                           | 45. ZnSO <sub>4</sub> _____                               |
| 21. Ca(NO <sub>3</sub> ) <sub>2</sub> _____            | 46. MnCO <sub>3</sub> _____                               |
| 22. BaSO <sub>4</sub> _____                            | 47. H <sub>2</sub> SO <sub>4</sub> _____                  |
| 23. BPO <sub>3</sub> _____                             | 48. Au <sub>2</sub> SO <sub>4</sub> _____                 |
| 24. LiNO <sub>3</sub> _____                            | 49. NH <sub>4</sub> NO <sub>3</sub> _____                 |
| 25. Al(NO <sub>3</sub> ) <sub>3</sub> _____            | 50. LiClO <sub>3</sub> _____                              |

# Naming Ionic Compounds With Polyatomic Ions Answers

**SMITH**



## **Naming Ionic Compounds With Polyatomic Ions Answers:**

*Survival Guide to General Chemistry* Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

*The Practice of Chemistry Study Guide & Solutions Manual* Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

*The Practice of Chemistry* Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POCPowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt) 832KB

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**Introduction to General, Organic, and Biochemistry** Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

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