5 • Chemical Bonding

LEWIS STRUCTURES

F	0	к	Al
of electrons =	# of electrons =	# of electrons =	# of electrons = _3
:F.	: Ö.	K-	-Al-
r- 7+1 of electrons = 8	02- 6+2 # of electrons = 6	K+ 1 - 1	A13+ 3-3
[:Ë:]	[:Ö:]²-	[KJ+	[Al]3+
2 7+7 of electrons = 14	H ₂ (+ (HF 1+7 # of electrons = 8	NH ₃ 5+3
:Ë:Ë:	H:H	H: Ë:	H: N: H
CH4 444 Fof electrons = 8	NF3 5+ 2-1 # of electrons = 26	SiF ₄ 4+ 28	C ₂ H ₆ 8+6 # of electrons = 14
H= C: H	·Ë · Ď · Ë ·	: Ë: S: : Ë:	H: C: C: H

MgH₂ 2+2

H : Mg : H

LiH |+ (

Li:H

H:AL:H

H

AIH₃

3+3

rl: H

H:B: H

H

BH₃

3+3

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