

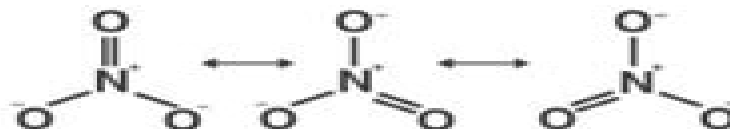
(i) Ozone molecule

(ii) Nitrate ion

Solution:



Resonating structure of Ozone



Nitrate resonating structure

49. Predict the shapes of the following molecules based on hybridisation.

BCl_3 , CH_4 , CO_2 , NH_3

Solution:

In compound BCl_3 , Boron has sp^2 -hybridisation and the shape is Triangular Planar.

In methane CH_4 , Carbon has sp^3 -hybridization and shape are Tetrahedral.

In carbon dioxide CO_2 , carbon has sp -hybridisation and shape is Linear.

In ammonia NH_3 , nitrogen has sp^3 -hybridisation and shape is Pyramidal.

50. All the C—O bonds in carbonate ion (CO_3^{2-}) are equal in length. Explain

Solution:

Carbonate ion is present in the form of a resonating structure. These structures are equivalent to nature. Resonance all 3 C-O bonds get a double character in one of the resonating structures.

51. What is meant by the term average bond enthalpy? Why is there a difference in bond enthalpy of O—H bond in ethanol ($\text{C}_2\text{H}_5\text{OH}$) and water?

Solution:

Similar bonds in a molecule do not have the same bond enthalpies. Mainly the term average bond enthalpy is used in polyatomic molecules. It is obtained by dividing bond dissociation enthalpy by the number of bonds broken. The bond enthalpy of OH bond is different in ethanol and water because of the difference in electronegativity of hydrogen and carbon. As electronegativity differs in hydrogen and oxygen is higher than that in carbon and oxygen, so the O-H bond in water has more bond enthalpy than in ethanol.

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