

Q16. Dinitrogen and dihydrogen react with each other to produce ammonia according to the following chemical equation:



- Calculate the mass of ammonia produced if 2.00×10^3 g of dinitrogen reacts with 1.00×10^3 g of dihydrogen.
- Will any of the two reactants remain unreacted?
- If yes, which one and what would be its mass?

Q17.

- Calculate the percentage of water of crystallisation in the sample of Mohr salt, $\text{FeSO}_4(\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$.
- Write the empirical formula of the compounds having the molecular formula:
 - C_6H_{12}
 - B_2H_6
 - N_2O_4
 - H_3PO_4
 - Fe_2O_3

Q18. State law of definite proportions. How can it be verified experimentally?

Q19.

- Calculate the weight of carbon monoxide having the same number of oxygen atoms as present in 88 g of carbon dioxide.
- Silver is a very precious metal and is used in jewellery. One-millionth of silver weighs 1.79×10^{-6} g. Calculate the atomic mass of silver.

Q20. A crystalline salt, on being rendered anhydrous, loses 45.6% of its weight. The composition of anhydrous salt is: Al = 10.5%, K = 15.1%, S = 2.8%, and O = 49.6%. Calculate the formula of anhydrous salt and crystalline salt.

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SA Adler



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