

**Assignment #2 – Compound Names and Formulas**  
**Multi-valent ions only**

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

1. copper (II) oxide CuO
2. mercury (I) oxide Hg<sub>2</sub>O
3. gold (III) chloride AuCl<sub>3</sub>
4. thallium (III) bromide TlBr<sub>3</sub>
5. bismuth (V) oxide Bi<sub>2</sub>O<sub>5</sub>
6. terbium (III) oxide Tb<sub>2</sub>O<sub>3</sub>
7. uranium (VI) oxide UO<sub>3</sub>
8. protactinium (IV) bromide PaBr<sub>4</sub>
9. cerium (III) oxide Ce<sub>2</sub>O<sub>3</sub>
10. arsenic (V) sulphide As<sub>2</sub>S<sub>5</sub>

11. manganese (III) oxide Mn<sub>2</sub>O<sub>3</sub>
12. vanadium (II) bromide VBr<sub>2</sub>
13. niobium(V) oxide Nb<sub>2</sub>O<sub>5</sub>
14. titanium (III) oxide Ti<sub>2</sub>O<sub>3</sub>
15. titanium (III) nitride TiN
16. iron (II) oxide FeO
17. cobalt (II) phosphide Co<sub>3</sub>P<sub>2</sub>
18. tin (II) oxide SnO
19. thulium (II) bromide TmBr<sub>2</sub>
20. copper (II) bromide CuBr<sub>2</sub>

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

1. SnCl<sub>4</sub> Tin (IV) Chloride
2. BiBr<sub>5</sub> Bismuth (V) Bromide
3. PoO<sub>2</sub> Polonium (IV) Oxide
4. PbI<sub>2</sub> Lead (II) Iodide
5. HgO Mercury (II) Oxide
6. HgCl Mercury (I) Chloride
7. Au<sub>2</sub>O<sub>3</sub> Gold (III) Oxide
8. FeCl<sub>2</sub> Iron (II) Chloride

9. PdF<sub>4</sub> Palladium (IV) Fluoride
10. Os<sub>2</sub>O<sub>3</sub> Osmium (III) Oxide
11. MoBr<sub>2</sub> Molybdenum (II) Bromide
12. VCl<sub>5</sub> Vanadium (V) Chloride
13. Mn<sub>2</sub>O<sub>3</sub> Manganese (III) Oxide
14. CoO Cobalt (II) Oxide
15. Np<sub>2</sub>O<sub>3</sub> Neptunium (III) Oxide
16. V<sub>2</sub>O<sub>5</sub> Vanadium (V) Oxide

# Naming Ionic Compounds Practice Answer Key

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*Chemistry in the Community (ChemCom)* American Chemical Society, 2011-06-17 Touted as the most successful NSF funded project published Chemistry in the Community ChemCom by the American Chemical Society ACS offers a meaningful and memorable chemistry program for all levels of high school students ChemCom covers traditional chemistry topics within the context of societal issues and real world scenarios Centered on decision making activities where students are responsible for generating data in an investigating analyzing that data and then applying their chemistry knowledge to solve the presented problem The text is intensively laboratory based with all 39 of the investigations integrated within the text not separate from the reading With the ChemCom program students learn more organic and biochemistry more environmental and industrial chemistry and more on the particulate nature of matter than other textbooks all within the relevance of solving problems that arise in everyday life Meticulously updated to meet the needs of today s teachers and students the new sixth edition of ChemCom adheres to the new science framework as well as the forthcoming next generation of science standards Incorporating advances in learning and cognitive sciences ChemCom s wide ranging coverage builds upon the concepts and principles found in the National Science Education Standards Correlations are available showing how closely aligned ChemCom is to these and other state standards

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A Self-study Guide to the Principles of Organic Chemistry Jiben Roy,2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix

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