

Half Modern Chemistry Review

CHAPTER 5: THE PERIODIC TABLE

The following pages contain the bulk (but not all) of the information for this chapter. I trust.

Focus on this material, but make sure to review other topics: reactions, bonding, molecules, etc.

If you study this document and UNDERSTAND what you should at least be able to DO after that.

Good items will be recall, examples, and/or application of this material. ~~Good~~

OUTLINE:

- Cells/mixes with periodic table/chemical/chemistry concepts. PC 12
- Communication/chemistry concepts to questions and period. PC 12
- 5.1: Describe the organization of the periodic table: distinguish between periods and groups. IT 41 42
- 5.2: Use the Periodic Table to obtain information about the properties (physical and chemical, correlation, energy, and atomic size) of the elements. IT 41 42
- 5.3: Identify common compounds, metalloids, and noble gases; identify the representative groups and/or other elements (groups 1, 2, 13, 14, 15, 16, 17, 18) IT 41 42
- 5.4: Compare periodic trends of atomic size, ionization energy, and electronegativity. IT 41 42

5.1: HISTORY OF THE PERIODIC TABLE

• Introductory

- periodic law = the law that states that the repeating chemical and physical properties of elements change periodically with the atomic numbers of the elements.
- periodic table = an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column.
- classification = a method of the periodic table of elements, which places elements, groups, ranges from 18 (groups 1-18) (transition).
- classification = one of the ways of classifying elements that are based on the periodic table (groups 1-18) (transition) (groups 1-18) are the periodic table.

• Chapter Highlights

- The periodic law states that the physical and chemical properties of the elements, and periodic functions of their atomic numbers.
- The periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family.
- The columns in the periodic table are referred to as groups.

5.2: ELECTRON CONFIGURATIONS AND THE PERIODIC TABLE

• Introductory

- atomic number = one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium).
- atomic number = one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium).
- transition elements = one of the elements that are not the same as the other elements in the periodic table.
- transition elements = one of the elements in the periodic table (groups 1-18) (transition) (groups 1-18) are the periodic table.
- transition = one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine) (halogens) (groups 1-18) (transition) (groups 1-18) are the periodic table.

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