

Selected Answers

Pre-Course Practice

DECIMALS (p. 20)	1.211	2.6795	3.15105	7.663
2.26304	11.729008	12.37	18.035	

FACTORS AND MULTIPLES (p. 302) 1, 1, 2, 3, 4, 6, 12
a. 1, 2, 3, 6, 9, 18, 27, 54 b. $2 \cdot 3^3$ c. 5 · 7 d. 1, 2, 4
11, 1, 2, 7, 14 13, 4 15, 3 17, 6 19, 2 21, 36
23, 42 25, 42 27, 900 29, 24 31, 60 33, 28 35, 54

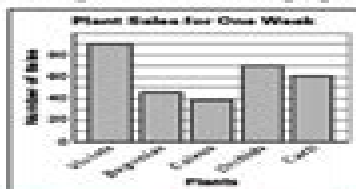
FRACTIONS	(p. 10)	1. $\frac{1}{10}$	2. $\frac{5}{10}$	3. $\frac{3}{10}$	4. $1\frac{1}{10}$	5. $\frac{9}{10}$
11. $\frac{19}{100}$	12. $\frac{3}{100}$	13. $\frac{9}{100}$	14. $\frac{1}{100}$	15. 6	16. $1\frac{1}{100}$	
17. $1\frac{1}{10}$	18. $1\frac{1}{10}$	19. $1\frac{1}{10}$	20. $1\frac{1}{10}$	21. $1\frac{1}{10}$	22. $1\frac{1}{10}$	

FRACTIONS, DECIMALS, AND PERCENTS			
$\frac{2}{100}$	2.038,	$\frac{19}{100}$	19.038,
$\frac{11}{100}$	11.13%,	$\frac{13}{100}$	13.160%,
$\frac{17}{100}$	17.06, 60%	$\frac{18}{100}$	18.063, 63%
$\frac{20}{100}$	20.064,	$\frac{21}{100}$	21.064,
$\frac{22}{100}$	22.064,	$\frac{23}{100}$	23.063,

COMPARING AND ORDERING NUMBERS (p. 68)											
1.	13.458	<	14.455	2.	-8.344	>	-8.434	3.	0.58	>	0.578
7.	1507	>	150	8.	$\frac{9}{10}$	>	$\frac{11}{10}$	9.	$\frac{1}{10}$	>	$\frac{1}{100}$
10.	1705	<	1705	11.	-0.205	<	-0.035	12.	-0.019	<	-0.013
13.	$\frac{4}{10}$	<	$\frac{1}{10}$	14.	$\frac{1}{10}$	<	$\frac{1}{10}$	15.	$\frac{1}{10}$	<	$\frac{1}{10}$
16.	$\frac{1}{10}$	<	$\frac{1}{10}$	17.	$\frac{1}{10}$	<	$\frac{1}{10}$	18.	$\frac{1}{10}$	<	$\frac{1}{10}$
19.	$\frac{1}{10}$	<	$\frac{1}{10}$	20.	$\frac{1}{10}$	<	$\frac{1}{10}$	21.	$\frac{1}{10}$	<	$\frac{1}{10}$
22.	$\frac{1}{10}$	<	$\frac{1}{10}$	23.	$\frac{1}{10}$	<	$\frac{1}{10}$	24.	$\frac{1}{10}$	<	$\frac{1}{10}$

PERIMETER, AREA, AND VOLUME (p. 101) 1. 10 m
a. 22.6 km b. 95 ft c. 3.92 in.² d. 39,304 ft²
e. 73.65 mm²

DATA DISPLAYS (p. 108) 1. Sample answer: 0 to 60 by five: 0, 10, 20, 30, 40, 50, 60 2. Sample answer: 0 to 25 by five: 0, 5, 10, 15, 20, 25
3. Sample answer: bar graph.



MEASURES OF CENTRAL TENDENCY (p. xxix) 1. 49; 5; 7
a. 52 1: 53: no mode

Chapter I

STUDY GUIDE 1. B 2. A 3. B 4. A

1.1 GUIDED PRACTICE (p. 4) 7. p minus 4, subtraction
8. 3 times x , multiplication 11. 1 12. $\frac{1}{11}$ 13. 54

1.1 PRACTICE AND APPLICATIONS (pp. 4–6) 21. 20
23. 2 25. 30 27. 9 29. 70 31. 6 33. 360 mi
35. 40 ft 37. 3,400 mi 39. 240 ft 41. 64 mi 43. 10 m²
45. 6 yd² 49. 4 h 51. 9.45 53. 15 57. $\frac{1}{8}$ 59. 13.9
61. 11,508 63. 53.55 65. 13,405

1.2 Super Practice 1a, 1b, 5, 6, 7, 8, 9, 11, 25

1.2 PRACTICE AND APPLICATIONS (pp. 12–14) 13, 23
 10. 9° 17. 5° 19. 52° 25 21. 16 23. 64 25. 1 27. 0
 29. 729 31. 32 33. 125 35. 371,293 37. 35,831,808
 39. 531,441 41. 29 43. 9 45. 20 47. 6 49. 15,625
 51. 100,000 53. 8 m³ 55. 2³, 8 cubic units 57. 4³
 64 cubic units 59. 18 61. 45 63. 9 71. 28 73. 3
 75. 9 77. 5 79. $\frac{10}{3}$ 81. $\frac{6}{7}$ 83. 9 85. 3 87. 7
 89, 91, and 93. Estimates may vary. 89. about 0.3; 0.27
 91. about 5; 4.764 93. about 6; 6.325

1.3 GUIDED PRACTICE p. 10 **2.50** **5.12** **7.17**
9.23 **11.4** **12.246** **15.3**

1.3 PRACTICE AND APPLICATIONS (pp. 18–21) 17. 34
 19. 1 21. 82 23. 300 25. 42 27. 11 29. 16 31. 43
 33. 14 35. 46 37. 3 39. $\frac{1}{2}$ 41. 128 43. 35(\$230 +
 \$300 + \$40 + \$15 + \$100 + \$200) = \$2000 45. $\frac{3}{4}$
 47. 2(\$7) + \$5 + 2(\$4) 49. 8 51. 162 53. 11 55. $\frac{3}{4}$
 57. 81 59. 900 61. composite; 1, 3, 9 63. composite;
 1, 2, 19, 38 65. composite; 1, 2, 5, 10, 25, 50 67. prime

GRADE 1 (p. 21)	1.12	2.14	3.32	4.9	5.5	6.16
7.6	8.54	9.216	10.200	11.2000	12.20	13.20
14.6 ³	14.4 ²	15.(3) ²	16.3 ²	17.(2x) ⁴	18.8 ²	
19.64 2 ²	20.2	21. $\frac{1}{2}$	22. $\frac{1}{2}$			

11. solution 13. not a solution 15. solution 17. not a solution 19. solution 21. not a solution 23. solution 25. solution

1.4 PRACTICE AND APPLICATIONS (pp. 27–28) 27. not a solution, 28. solution, 29. solution, 30. solution, 31. solution, 32. solution, 33. solution, 34. solution, 35. solution, 36. solution, 37. 3, 38. 9, 39. 41, 21, 43, 2, 45, 5, 47, 6, 51. solution, 52. not a solution, 53. solution, 54. solution, 55. solution, 56. solution, 57. 3.4 boxes or more, 58. 7, 2, 1, 59. 16, 60. 16, 61. 7, 2, 62. 7, 2, 63. 7, 2, 64. 7, 2, 65. 7, 2, 66. 7, 2, 67. 7, 2, 68. 7, 2, 69. 7, 2, 70. 12, 71. 7, 72. 7, 73. 7, 74. 7, 75. 7, 76. 7, 77. 7, 78. 7, 79. 7, 80. 7, 81. 7, 82. 7, 83. 7, 84. 7, 85. 7, 86. 7, 87. 7, 88. 7, 89. 7, 90. 7, 91. 7, 92. 7, 93. 7, 94. 7, 95. 7, 96. 7, 97. 7, 98. 7, 99. 7, 100. 7.

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aug 13 2021 pi artworks founded in 1998 pi artworks focuses on the meeting of perspectives of the art market through its mixture of turkish and international exhibitions it is widely recognised as one of the leading contemporary art galleries in istanbul since its beginnings the gallery has grown under the direction of yesim turanli who was also one

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web chemistry chemistry questions and answers lab 16 lewis dot structures molecular models vsepr theory valence bond theory and polarity name date formula number of lewis dot electrons formula molecular electron geometry drawing of structure hybridization geometry hinh trigonal 8 tetrahedral sp^3 pyramidal nh_3 h fb f

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web problem 4 2 7 4 2 7 the arrangement of atoms in several biologically important molecules is given here complete the lewis structures of these molecules by adding multiple bonds and lone pairs do not add any more atoms a the amino acid serine b urea c pyruvic acid d uracil e carbonic acid

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web introduction chemical structures there are a number of different ways to represent molecules as illustrated for nh_3 in figure 1 a lewis structure figure 1a is meant to represent chemical bonding and the distribution of electrons

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web nov 17 2023 introduction lewis structure is basically a graphic representation of the electron distribution around an atom the major reason why learning lewis dot structure is important is that it helps in predicting the number and type of bonds which can be formed around an atom it also helps in predicting the geometry of the molecule

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structure regions of e density electronic group geometry name hybridization sp sp² sp³ of bonding regions of lone pairs shape of the molecule drawing and name polar or nonpolar p or np h₂o

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web nov 14 2023 show all resonance structures possible use formal charges to pick the best form if necessary the following molecule has two lewis dot formulas write one that uses 8 electrons about the central atom and one that expands the natat

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web lewis structures a lewis structure is a representation of covalent molecules or polyatomic ions where all the valence electrons are shown distributed about the bonded atoms as either shared electron pairs bond pairs or unshared electron pairs lone pairs a shared pair of electrons is represented as a short line a single bond

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web 2019 0501140012 american native medicine 1 2019 0612171451 orpic company 1 biome viewer student ws copy of lab 1 effect environment se the perodic properties of the main group to understand theory behind lewis structure and valence shell electron pair repulsion theory to experience molecule s geometrical shapes by the model that has

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web mar 27 2023 a the lewis structure for water is represented by the given option a in this structure the oxygen atom is in the center and is bonded to two hydrogen atoms the oxygen atom has two lone pairs of electrons this structure follows the octet rule where each atom has a full valence shell with eight electrons

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web introduction the physical and chemical behavior of compounds depends on the types of bonds that exist within the molecule and on the overall shape of the molecules the branch of chemistry that investigates how molecules are put together is

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web lewis structures represent the valence electrons present in the molecule as lone pairs unshared electrons drawn as dots or bonds shared electrons drawn as lines procedure for drawing lewis structures add up the total number of valence electrons that each atom contributes to the molecule polyatomic ion

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web the pcl 3 lewis structure has the typical case of phosphorus p in the center with 3 bonds to 3 other atoms phosphorus is from the same column as nitrogen in the periodic table meaning that p and n generally have the same bonding structure note the lone pair dots without bonds on top of p just like for n in the previous example for nh 3 chlorine cl

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web step 1 figure out how many electrons the molecule must have based on the number of valence electrons in each atom when drawing the structure of an ion be sure to add subtract electrons to account for the charge step 2 connect the atoms to each other with single bonds to form a skeleton structure

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web molecular models and 3d printing activity lewis dot structures and molecule geometries worksheet answer key 6 10 co 3 2 total number of valence electrons 24 cad engineered 3d sketch model show dipole arrows lewis structure is there a polar bond in this molecule yes or no vsepr shape name trigonal planar bond angles 120

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web jun 8 2022 lewis structure in chemistry lewis structures also called electron dot structures or electron dot diagrams are diagrams that show the bonding between atoms of a molecule and the lone pairs of

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web this exercise required us to take the formulas in the data table and calculate the number of valence electrons using the periodic table draw its lewis structure on paper build the molecular model and record a photo list the number of bonds and lone pairs surrounding the central atom and lastly identify the structure geometry

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