

ANSWERS

Section III: Problems (Answers to odds in back of the text book)

Practice Problems:

Chem Practice Problems:

- Ammonia is a strong household cleaner. What is the atomic mass of one molecule?
17.04amu
- The double replacement ...
- A bottle of lead II sulfate contains 158.1g of the compound. How many moles of the compound are present?
0.5213 mol PbSO₄
- Find the
- Determine the number of formula units that are in 1.48 mol NaF
8.91 x 10²³ units NaOH
- How many formula ...
- How many atoms are in 1.25 mol of oxygen gas?
1.27 x 10²⁴ atoms O
- If you burned 6.10 x 10²⁴ ...
- A container with a volume of 101L contains how many moles of argon gas at STP?
4.51 mol Ar
- Find the percent composition ...
- A sample of an unknown compound with a mass of 0.847g has the following composition 50.51% F & 49.49% Fe (m/m). When the compound is decomposed into its elements, what is the contributing mass of each? Write the decomposition equation after reassembling the empirical formula.
.428g F & .419g Fe

.0225mol F = 3 & .00751 mol Fe = 1; 2FeF₃ → 2Fe + 3F₂
.00751mol .00751 mol
- Find the molecular formula ...
- A 50.51g sample of made of phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39g of phosphorus atoms were produced.
 - What is the empirical formula of the compound?
$$\frac{11.39}{30.97} = \underline{\underline{.368\text{mol P}}} \text{ \& } \frac{(50.51\text{g} - 11.39\text{g})}{35.45\text{g}} = \underline{\underline{.551\text{mol Cl}}}$$
$$\underline{\underline{.368\text{ mol}}}$$

= 1 P & 1.5 Cl (double to make whole #s) P₂Cl₃

General Chemistry Homework Guide

ML Morrison



General Chemistry Homework Guide:

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