ACS Gen Chem 1 Exam

Ion - A species of an element in which the number of electrons does not equal the number of protons.

Isotope - Each isotope of an atom has a different number of neutrons.

Isoelectric - Two atoms with the same charge.

Isotopic - Two atoms with the same number of neutrons.

Transition Metal Orbitals - The 4s orbital is filled before the 3d orbital, but listed after.

Paramagnetic - Atoms that possess a permanent magnetic charge due to the presence of unpaired electrons.

Electropositive - Tending to lose electrons and form positive ions.

Tetrahedral Bond Angle - 109.5°

Resonance Structures - Two equally correct arrangements of one molecule's electrons; generally the molecule exists as an average of the two.

Polar Molecule - A molecule that is mostly positive on one side and mostly negative on the other.

sp³ Orbital - An orbital composed of one s orbital and three p orbitals; can hold eight total electrons. Tetrahedral.

Trigonal Planar - The way that three objects will arrange themselves around a central atom. Non-polar.

Free Radicals - Compounds with an odd number of electrons in their Lewis Structure. Eg. NO₂

Incomplete Octets - Occur in elements that can form stable bonds with less than eight electrons. Eg. BF₃

Expanded Octets - Occur in the third row of the periodic table and beyond in elements that can form stable bonds with more than eight electrons. Eg. XeF₂

Hydrogen Bonding - When hydrogen is attracted to electronegative atoms.

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Lei Huang

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Table of Contents General Chem 1 Exam Study Guide

- 1. Understanding the eBook General Chem 1 Exam Study Guide
 - The Rise of Digital Reading General Chem 1 Exam Study Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying General Chem 1 Exam Study Guide
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chem 1 Exam Study Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from General Chem 1 Exam Study Guide
 - Personalized Recommendations
 - General Chem 1 Exam Study Guide User Reviews and Ratings
 - General Chem 1 Exam Study Guide and Bestseller Lists

- 5. Accessing General Chem 1 Exam Study Guide Free and Paid eBooks
 - o General Chem 1 Exam Study Guide Public Domain eBooks
 - General Chem 1 Exam Study Guide eBook Subscription Services
 - General Chem 1 Exam Study Guide Budget-Friendly Options
- 6. Navigating General Chem 1 Exam Study Guide eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - General Chem 1 Exam Study Guide Compatibility with Devices
 - General Chem 1 Exam Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chem 1 Exam Study Guide
 - Highlighting and Note-Taking General Chem 1 Exam Study Guide
 - Interactive Elements General Chem 1 Exam Study Guide
- 8. Staying Engaged with General Chem 1 Exam Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chem 1 Exam Study Guide
- 9. Balancing eBooks and Physical Books General Chem 1 Exam Study Guide
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection General Chem 1 Exam Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine General Chem 1 Exam Study Guide
 - Setting Reading Goals General Chem 1 Exam Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chem 1 Exam Study Guide
 - Fact-Checking eBook Content of General Chem 1 Exam Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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