Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion	Polyatomic Anion Name
HCIO.	Chloric acid	H,O-	C/O3-1	Chlorate
H,SO,	Sulfarous acid	2450	503-2	Sulfite
H,SO,	Sulfuric acid	2150	500-3	Sulfate
H.PO,	Phosphorous acid	3 H,O-	P03-3	Phosphite
H.PO.	Phosphoric acid	3 H50+	PO,5-	Phosphate
HNO,	Nitric acid	H30+	NO."	Nitnate
HNO,	Nisrous acid	H.O+		Nitrite
H,CO.	Carbonic acid	2 H30+	CO,-2	Carbonate

- 6. Look at the formulas of the ternary acids in Model 2.
 - at. How are ternary acids different from binary acids in their structure?

Ternory acids contain three elements and they also contain the element oxygen.

6. What number do you think the prefix "ter-" refers to?

"te" refer to three

- 7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.
- 8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is /C.

- Polyatomic anion ending is "-ite" → acid name ending is aus
- 9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when
- naming HClO? C/O3' has the name of Chlorate. "ate" must be replaced with "ic. If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl, not HClO3.

 The prefix "hydro" is not used in the name when an acid contain more than two elements and are of the elements is axygen.

 The prefix "hydro" is not used in the name when an acid contain more than two elements and are of the elements is axygen.

 It Predict the formula for chlorous acid.

12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."

HISO.



H,CO.

HNO.

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