

Summary NACE CP2 Study Guide 2023

Complete Solutions

oxidation - ANSWER ✓✓loss of electrons, positive charge atoms (anode)

reduction - ANSWER ✓✓gain of electrons, negative charged atoms (cathode)

forms of corrosion - ANSWER ✓✓uniform, crevice, galvanic (dissimilar metals) if two metals have to be couple then best to couple metals close together in the galvanic series, pitting, intergranular, selective leaching (dezincification, graphitization), velocity phenomena, environmental cracking (corrosion fatigue, hydrogen embrittlement, SCC (Carbonates/bicarbonates)

High PH - ANSWER ✓✓Low hydrogen, alkaline, easy polarization

Low PH - ANSWER ✓✓high hydrogen, acidic, harder to polarize

High oxygen/High Hydrogen concentration - ANSWER ✓✓harder to polarize, work at the cathode only, called cathodic depolarizers.

factors affecting corrosion - ANSWER ✓✓A. anything that affects polarization. (temp. etc...)

B. concentration cells (oxygen, temp, etc.) In an oxygen concentration cell, what is the anode? (i.e. the area with more or less oxygen?) In a metal ion cell, which area is the anode (i.e. the area with the greater or lower concentration)?

C. Anaerobic bacteria (MIC) Absence of oxygen, depolarizer

polarization - ANSWER ✓✓a. corrosion current reduces with polarization

b. polarization increases with CP current applied

c. Occurs at surface of both anode and cathode.

shielding - ANSWER ✓✓a. shorted casings

b. proximity of other structures

Nace Cpstudy Guide

Didier Musso



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