# OCR: Chemistry – Unit 2 – Chains, Energy and Resources

Note: For all Mechanisms and Reactions - There is an attached sheet at the end showing all the needed reactions in more detail.

THERE IS ALSO A LIST OF ALL EQUIATIONS AND REACTION NEEDED FOR THIS SYLLABUS ATTACHED IN A SEPARATE FILE.

(a) Interpret and use the terms:

 Empirical formula as the simplest whole number ratio of atoms of each element present in a compound.

 (ii) molecular formula as the actual number of atoms of each element in a molecule,

(iii) general formula as the simplest algebraic formula of a member of a homologous series

Alkane	C <sub>3</sub> H <sub>25+2</sub>	
Alkene	C <sub>n</sub> H <sub>2n</sub>	
Alcohol	C <sub>2</sub> H <sub>2/22</sub> OH	
Cycloalkane	C <sub>n</sub> H <sub>2n</sub>	
Cycloalkene	C <sub>n</sub> H <sub>2n-2</sub>	
Haloalkane	C.H.m.X	=======================================

(iv) structural formula as the minimal detail that shows the arrangement of atoms in a molecule

E.g. Butane:

СНЗСН2СН2СН3 or CH3(CH2)2CH3,

(v) displayed formula as the relative positioning of atoms and the bonds between them, ie for ethanol:

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Yan Bai

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